

Influence of Drying on the Characteristics of Zinc Oxide Nanoparticles

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The recent growth in the field of porous and nanometric materials prepared by non-conventional processes has stimulated the search of new applications of ZnO nanoparticulate. Zinc oxide is an interesting semiconductor material due to its application on solar cells, gas sensors, ceramics, catalysts, cosmetics and varistors. In this work, the precipitation method was used followed by controlled and freezing drying processes. The materials obtained were thermally treated at various temperatures. The influence of temperature on structural, textural, and morphological properties of the materials was studied by powder X-ray diffraction, infrared spectroscopy, scanning electron microscopy, nitrogen adsorption, and thermal analysis. The characteristics of both materials were compared.

Keywords: zinc oxide, drying process, nanoparticle

I. INTRODUCTION

Zinc oxide is a versatile material due to properties such as: pyroelectricity, semiconducting, piezoelectricity, luminescence, and catalytic activity [1–2]. Its optical band gap, chemical and thermal stabilities are also very important characteristic. These properties are essential for electronic and photonic industry due to the wide range of technological applications in devices including acoustic surface, wave filters, sensors, transparent conductors, light-emitting diodes, laser diodes, solar energy conversion [3–5], electro-acoustic transducers, photovoltaic devices, catalysts and varistors [6–8]. Zinc oxide powders are also a commercially important material utilized in paints and rubber [9].

The properties of ZnO depend mainly of the crystallite size and morphology [10]. With the objective to prepare ZnO nanostructures with desire properties, some methods are used, such as the sol-gel process [11], hydrothermal route, chemical vapor deposition, homogeneous precipitation and electro-deposition. Farley et al produced ZnO from zinc acetate as precursor, getting a high quality semiconductor oxide for various technical applications [12]. The hydrothermal process was used to obtain different structures of ZnO [13]. ZnO nanoparticulate was prepared by solid state reaction at low temperature of calcination (500 °C) through thermal decomposition, [14]. The evolution of the microstructure during the formation of ZnO powder particles obtained by the freeze-drying method was investigated by Marinkovic et al [15]. These syntheses allowed to the preparation of ZnO with a variety of morphologies, including nanoparticles, nanowire, nanorod, nanoribbon, nanoplate, nanotube, multipods, etc [16–18].

In this work the precipitation process was used to obtain ZnO particles by reaction between aqueous solutions of $Zn(NO_3)_2$ and Na_2CO_3 . The textural, microstructural, and morphological characteristics were investigated to establish the correlation between the powder characteristics with the drying type and the calcination temperature.

II. EXPERIMENTAL

Zinc nitrate and sodium carbonate were used as precursors of the ZnO particles. In a typical synthesis, $Zn(NO_3)_2 \cdot 6H_2O$ were dissolved in deionized water in the molar ratio of 1:5, and mixed for homogenization during 1 h. Na_2CO_3 were also dissolved in deionized water in the molar ratio of 7:10, and mixing with the zinc solution leading to the formation of a white precipitate. The precipitate was washed with deionized water to remove impurities. The obtained powder was separated in two aliquots to be submitted to drying processes. One aliquot (S1) was dried in an oven at 110°C for 48 h and the other (S2) was immersed in liquid nitrogen to freezing and then was lyophilized by 78 h at -40 °C and vacuum of 0.7 Pa. Both the aliquots were calcined at different temperatures between 200 and 1000°C for 2 hours.

Simultaneous thermogravimetric and differential thermal analysis (TG–DTA) measurements were performed (TAInstrumentSDT2960). The samples were heated from room temperature up to 1100°C at 10°C.min⁻¹ under airflow.

All samples were analyzed by X-ray diffractometry (Rigaku, Geigerflex3034) with CuK α radiation, 40 kV and 30 mA, time constant of 0.5 s and crystal graphite monochromator. Crystallite size was determined by Scherrer equation ($D = 0.9\lambda/\beta\cos\theta$, where D is the crystallite diameter, λ is the radiation wavelength and θ the incidence angle) [17]. The value of β was determined from the experimental integral peak width with silicon as a standard. Values were corrected for instrumental broadening.

Infrared spectroscopy (Perkin–Elmer, 2838) was used to evaluate the chemical modifications of ZnO powders heated at different temperatures. The morphology of the composites was evaluated by scanning electron microscopy (JEOL–JSM840A). The textural characteristics of the samples were determined through nitrogen gas adsorption (Autosorb - Quantachrome Nova 1200) at liquid nitrogen temperature. Nitrogen gas was used with a 22-point adsorption-desorption cycle. The samples were outgassed at 200°C for 3 h before each analysis. Specific surface areas and total pore volumes were obtained by the application of the Brunauer-Emmett-Teller (BET) equation and the BJH method. The particle size of the powders was evaluated using the equation: $d(nm) = 6/S \cdot \rho$, where 6 is the form factor for spherical or cubic particles, S is the superficial area in m² g⁻¹ and ρ is the

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true density in g cm^{-3} .

III. RESULTS AND DISCUSSION

TG curve for the S1 sample (Fig.1a) shows three stages of weight loss accompanied by two endothermic and two exothermic peaks in DTA curve. The first endothermic event between 100 and 200°C can be associated to removal of surface water of the sample with weight loss of 18%. The weight lost of 2% in the range of 210-380°C, accompanied by one endothermic and one exothermic peak, and of 2 % in the range of 410-800°C is due to the nitrate and CO₃ decomposition. The broad and intense exothermic peak could be associated with the formation of the pure ZnO phase. The obtained residue is around 78% (pure ZnO phase).

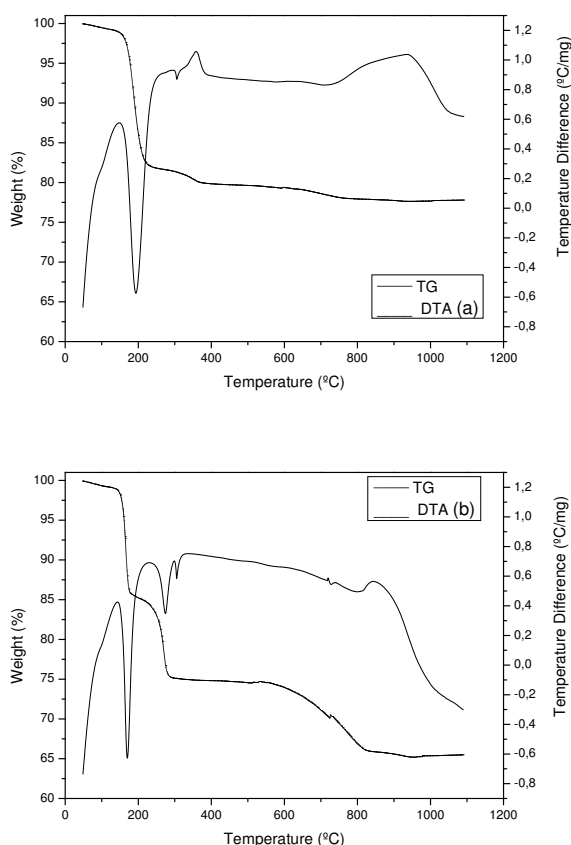
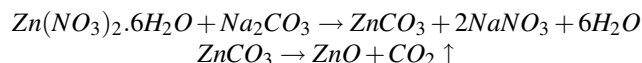


FIG. 1: TG/DTA curves of ZnO previously dried at 110°C, (a) controlled and (b) freeze drying.

The S2 sample (Fig.1b) lost 14% weight in the range of 50-180°C due to surface water removal related to an endothermic event. The other two endothermic events between 180-250°C and 580-820°C were attributed to loss of sodium nitrate (11 %) and carbon dioxide (10 %). These events are accompanied by endothermic peaks in DTA curves. A broad exothermic event appears between 800 and 930°C, probably due to the formation of the pure ZnO phase. In this case, the residue obtained is around 65 %. We can describe the ZnO formation

with the reaction below:



Figures 2a and 2b show XRD patterns of S1 and S2 samples calcined at various temperatures. The characteristic peaks of the ZnO phase increase in intensity above 200°C with the calcination temperature. The S1 sample presented peaks characteristics of ZnO phase at 200°C and by-products such as NaNO₃ and ZnCO₃ (4 %), which decomposed gradually with increasing in the temperature. The pure phase of ZnO was formed below 800°C.

The S2 sample shows peaks of ZnO phase with 21% of by-products such as ZnCO₃ and NaNO₃ between 200 and 600°C, which decomposed with the increase in the temperature, originating pure phase of ZnO above 800°C.

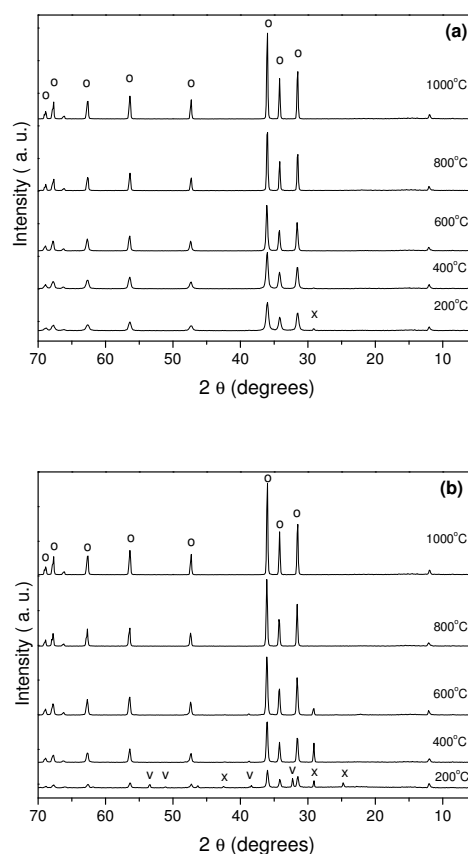


FIG. 2: X-ray diffraction patterns of o = ZnO, x = NaNO₃ and v = ZnCO₃ powders treated at different temperatures a) prepared by controlled drying (S1) and b) prepared by freeze drying (S2).

Figure 3 shows infrared spectra of the ZnO samples (S1 and S2) treated at various temperatures. The spectra show bands (3500 cm^{-1}) assigned to OH⁻ stretching vibrations of crystalline and adsorbed H₂O. Two bands were observed at 1512 and 1365 cm^{-1} , corresponding to ν_3 frequency of carbonate groups (C – O and C = O). In the spectra of S1 sample, these bands disappear after treatment at 600°C and in S2 spectra after 800°C. Bands at 1050 cm^{-1} , 835 cm^{-1} and 703 cm^{-1} relate to ν_1 , ν_2 , and ν_4 frequencies of the carbonate

appear in S1 spectra only for treatment at 110°C and in S2 spectra until treatment at 800°C. Bands with small intensity characteristics of nitrate groups appear at 1630 cm⁻¹. These bands disappear only above heating at 800°C. The S1 spectra show a broad band between 550 and 465 cm⁻¹ with shoulder shape, characteristic of hexagonal Zn-O phase. With the increasing in the calcination temperature, the shoulder shape changes and a unique broad band is formed. The S2 spectra also show a broad band, without the shoulder shape, more intense than the S1 sample ones. According to literature [18], the shape of the IR bands can be due to the geometry and size of the particles, and/or the aggregate formation.

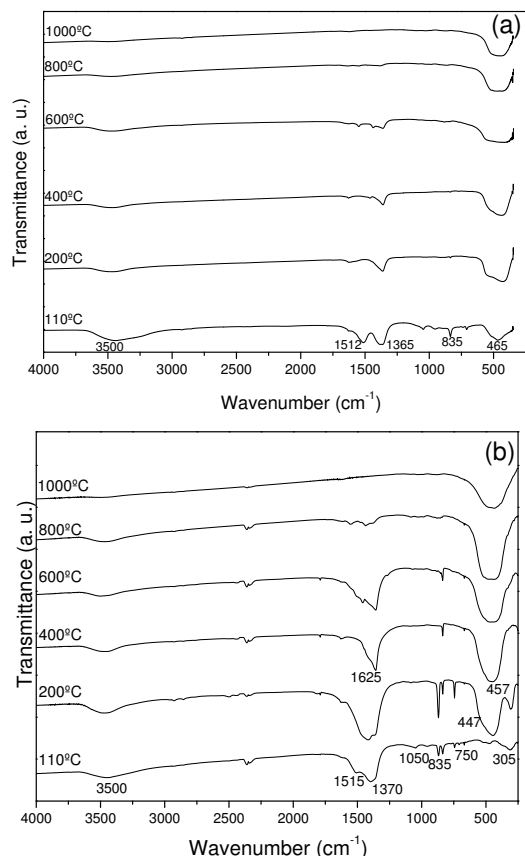


FIG. 3: Infrared spectra of ZnO powders treated at different temperatures a) prepared by controlled drying (S1) and b) prepared by freeze drying (S2).

The IR spectra show that the S2 sample has more impurities than the S1 sample, according to XRD patterns and TG/DTA curves, confirming the difference between the powders produced in both the process.

Figure 4 shows the morphology of the ZnO. The images obtained by SEM of the S1 samples treated at 110°C (Fig.4a) show plate-like nanoparticles forming agglomerates. After treatment at 800°C, these particles (Fig.4b) grow and densify, reaching an average particle size of $0.550 \pm 0.050 \mu\text{m}$. The S2 images (Fig. 4c) also shows plate-like nanoparticles, which grow after treatment at 800°C (Fig.4d), reaching an average particle size of $1.1 \pm 0.1 \mu\text{m}$. The samples showed significant morphological differences.

Table 1 shows the textural characteristics of ZnO samples

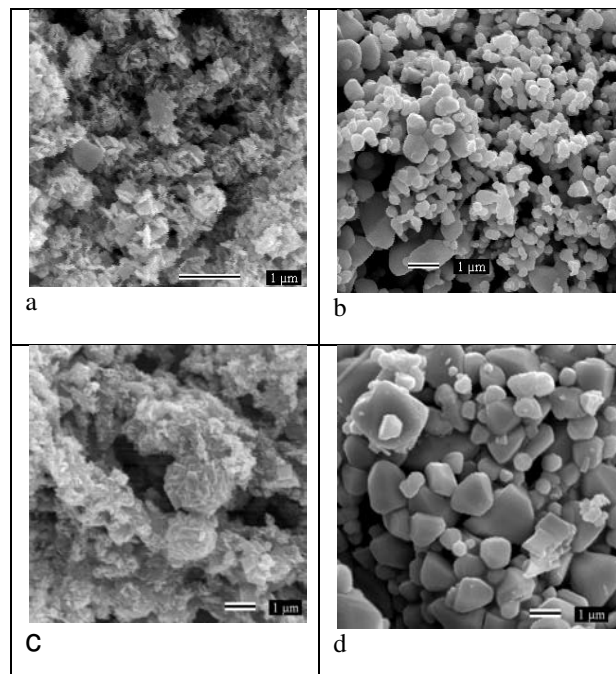


FIG. 4: Images of ZnO powders treated at different temperatures a) S1 - 100°C, b) S1 - 800°C, c) S2 - 100°C and d) S2 - 800°C.

Table 1 – Textural and morphological characteristics of the ZnO nanoparticles.

Temperature (°C)	SSA (m ² g ⁻¹)	True Density (g.cm ⁻³)	Total pore volume (10 ⁻³ cm ³ g ⁻¹)	Average particle size (nm)	Average crystallite size by XRD(nm)
Controlled drying					
110	39.1	3.8	48	40	12
200	27.2	3.9	35	57	29
400	17.1	4.2	23	84	35
600	7.2	4.5	10	185	45
800	2.1	4.8	2	595	64
Freeze drying					
As-prepared	20.1	2.5	31	119	46
200	7.2	3.0	8	278	35
400	2.0	3.3	1	909	66
600	1.5	3.8	2	1051	82
800	1.0	5.0	2	1200	91

treated at various temperatures. S1 samples had specific surface area, porosity and true density larger than S2 samples. We also can observe that the average particle size measured by BET method is according to the results obtained by SEM.

Figure 5 shows the adsorption-desorption isotherms for samples S1 treated at different temperatures. The material has characteristic of no-porous nanoparticulate material only until heating temperatures of 600°C. (type II isotherm, according to BDDT [19] classification). The hysteresis type H3 is associated with a closure of the hysteresis loop due to the so-called tensile strength effect, provoked by particles plate-like, according to the SEM images. This characteristic disappears with the increase in the heating temperature due to the growth and densification of the particles.

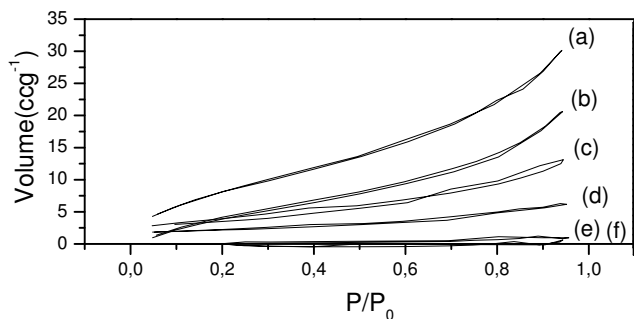


FIG. 5: Adsorption-desorption isotherms of ZnO nanoparticles (S1) treated at different temperatures: (a) 110°C, (b) 200°C, (c) 400°C, (d) 600°C, (e) 800°C and (f) 1000°C.

The freeze drying materials showed irregular isotherm curves (not shown), with C constant negative, characteristic of small interaction of the nitrogen with the material. We observed that the nitrogen did not wet the ZnO surface, showing

that this kind of drying change significantly the surface sample, confirming the IR results.

IV. CONCLUSION

The as-prepared zinc oxide obtained by controlled temperature (S1) was more pure than ZnO powder obtained by freeze drying (S2). These impurities were eliminated with the increasing in the calcination temperature but influenced drastically the morphological and textural characteristics of the final product. All the samples calcined at 800°C had the same structure but with different particle size and surface characteristic.

V. ACKNOWLEDGEMENT

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