

Time domain NMR evaluation of poly(vinyl alcohol) xerogels

Elton Jorge da Rocha Rodrigues¹, Maxwell de Paula Cavalcante¹ and Maria Inês Bruno Tavares^{1*}

¹*Centro de Tecnologia, Instituto de Macromoléculas Professora Eloisa Mano – IMA, Universidade Federal do Rio de Janeiro – UFRJ, Rio de Janeiro, RJ, Brazil*

*mibt@ima.ufrj.br

Abstract

Poly(vinyl alcohol) (PVA)-based chemically cross-linked xerogels, both neat and loaded with nanoparticulate hydrophilic silica (SiO₂), were obtained and characterized mainly through time domain NMR experiments (TD-NMR). Fourier-transform infrared (FT-IR) and wide angle X-ray diffraction (WAXD) analyses were employed as secondary methods. TD-NMR, through the interpretation of the spin-lattice relaxation constant values and related information, showed both cross-linking and nanoparticle influences on PVA matrix. SiO₂ does not interact chemically with the PVA chains, but has effect on its molecular mobility, as investigated via TD-NMR. Apparent energy of activation, spin-lattice time constant and size of spin domains in the sample have almost linear dependence with the degree of cross-linking of the PVA and are affected by the addition of SiO₂. These three parameters were derived from a single set of TD-NMR experiments, which demonstrates the versatility of the technique for characterization of inorganic-organic hybrid xerogels, an important class of materials.

Keywords: *inorganic-organic hybrid, polymer gels, relaxometry, low-field nuclear magnetic resonance, time domain nuclear magnetic resonance.*

1. Introduction

Time domain nuclear magnetic resonance (TD-NMR) – also known as relaxometry, wide-line or low-field nuclear magnetic resonance (LF-NMR) – is a rapid and inexpensive way to characterize polymeric materials. It is based on the measurement of the time necessary for relaxation to equilibrium of a magnetized atomic nucleus in a nuclear magnetic resonance experiment^[1-3]. This process is related to the molecular dynamics of the sample and can reveal many factors acting on the polymer matrix, such as variations in crystallinity, blending of two or more polymers, presence of plasticizing agents or fillers, cross-linking and others^[1,4-8]. This versatile tool has been applied in various fields, such as determination of physico-chemical parameters of crude oil samples, polymer-filler interactions in rubber materials, and reaction monitoring and quality assurance in industrial processes, to cite a few examples^[9-12].

Polymeric gels are networks containing physically and/or chemically cross-linked segments which can halt the long distance rearrangements between the polymer chains that are responsible for the material's dissolution and plastic deformation^[13]. When polymeric gels are not filled with fluids, they are dubbed xerogels^[14]. This class of polymeric material has many applications, such as in tissue engineering, drug delivery, prosthetics and membranes applied in separation processes^[14-16].

Various efforts have been made to characterize polymeric gels through TD-NMR^[15,17-22], but the majority of the reported methods involve soaking the sample in a solvent^[14,19,23,24], with the analysis being conducted under strict temperature conditions^[18,20,21], or limiting the chemical nature of the polymeric matrix to elastomers^[6,12,18-21].

In this work, we obtained xerogels based on chemically cross-linked poly(vinyl alcohol) (PVA) and PVA loaded with nanoparticulate hydrophilic silica (SiO₂). These materials were characterized by LF-NMR techniques and supporting spectroscopic characterization methods, such as Fourier-transform infrared (FT-IR) and wide angle X-ray diffraction (WAXD). Our aim was to use the information obtained mainly through TD-NMR to qualify and quantify the changes caused by cross-linking and inorganic filling of the PVA matrix.

2. Experimental

2.1 Materials

PVA (99+% hydrolyzed), M_w : 89,000-98,000 g.mol⁻¹ and glutaraldehyde (GA) (25% m/m in H₂O) were supplied by Sigma-Aldrich. Hydrochloric acid (HCl) (37% w/w) was obtained from Vetec, Brazil. SiO₂, composed of 12 nm radius particles, was supplied by Degussa. Deionized water was employed during all necessary steps. All reagents were used as received.

2.2 Sample preparation

Six 40 mL aliquots of a previously prepared 10% (by weight) aqueous PVA solution were used to obtain the samples, three for neat PVA xerogels and three for PVA/SiO₂ xerogels. To each of the PVA solutions, a sufficient volume of a 1 mol.L⁻¹ HCl solution was added to bring the pH of the system to 2±0.1. To each of the PVA/SiO₂ solutions, 1% (by weight of polymer) of SiO₂ was added to the PVA solution and left under stirring until all the nanoparticles

were homogeneously suspended. Then, a volume of GA enough to induce 1%, 5% and 10% of nominal cross-linking, based on the molar mass of PVA's repeating unit, was added to the acidified solutions. Finally, the contents of the six aliquots were poured into polystyrene Petri dishes, covered and left in a fume hood at room temperature for five days to allow completion of the chemical cross-linking reaction. After this period, the gels were thoroughly washed with deionized water and put in a vacuum oven at 45 °C for 2 days. Additionally, for comparison, a film of neat PVA and one of PVA/SiO₂ were produced following the same procedure, except for the acidification and the addition of GA steps. All the materials were stored in a desiccator under vacuum until needed for tests. Table 1 reports the naming scheme of the gels and polymeric films.

2.3 FT-IR analysis

The analysis was conducted on 1-mm thick films to investigate changes caused by the cross-linking reaction to the polymeric matrix, as well as interactions between PVA and SiO₂. The test was run in a Varian Excalibur FT-IR spectrophotometer, with a zinc selenide crystal for ATR measurements, having the following acquisition parameters: 100 scans, 4 cm⁻¹ spectral resolution, sampling between 600 and 4000 cm⁻¹.

2.4 WAXD analysis

Films similar to those employed for FT-IR measurements were subjected to WAXD analysis to verify changes in the molecular order of the PVA, using a Rigaku Miniflex diffractometer, operating with Cu K α band (1,5418 Å wavelength) radiation. The 2 θ angle was varied between 5° and 50°, with sweep speed of 0.05°·s⁻¹.

2.5 TD-NMR analysis

Relaxometry was employed to assess the samples' molecular dynamics through spin-lattice relaxation processes. All spectra were obtained in a Maran Ultra spectrometer, with an 18 mm magnet bore, operating at 0.54 T (23.4 MHz for ¹H) and equipped with a heating/cooling module. The samples were shredded in a mill for better conditioning inside the glass tube used as sample holder. The masses of each sample were measured to be similar to each other, so that the results could be better compared.

Table 1. Sample names.

Name	Nominal cross-linking	Percentage of inorganic nanoparticles
PVA	0	0
PVA/SiO ₂	0	1% (m/m)
P1	1% (m/m)	0
P1A	1% (m/m)	1% (m/m)
P5	5% (m/m)	0
P5A	5% (m/m)	1% (m/m)
P10	10% (m/m)	0
P10A	10% (m/m)	1% (m/m)

2.5.1 Spin-lattice relaxation constant determination

The inversion-recovery pulse sequence was chosen to extract the samples' spin-lattice relaxation time constant, T₁. The data were processed in commercial software packages, namely WinFit version 2.4.0.0 and Origin version 8.5. Table 2 shows the analytic parameters used to measure the samples' relaxation times.

The T₁ values displayed are the average between the results from two runs. This mean value was extracted from fitting the experimental points using Equation 1.

$$M(\tau) = A_0 + A_1 e^{\left(-\frac{\tau}{T_1}\right)} \quad (1)$$

where $M(\tau)$ is the magnetization as a function of time between 180° and 90° pulses; A₀ is the DC offset; A₁ is a factor proportional to the number of relaxing proton nuclei in each interval; τ is the time interval between 180° and 90° pulses; and T₁ is the spin-lattice relaxation time constant. Four different temperatures – 35, 45, 55 and 65 (±0.1) °C – were employed during the measurements, with the sample being left to equilibrate for 5 minutes at each step.

2.5.2 Spin diffusion path

T₁ relaxation times of solid polymers are not determined solely by dynamic processes. Below the glass transition temperature (T_g), spin diffusion also contributes to the relaxation time value. This diffusion can be modeled as Fickian diffusion and can be useful to understand phase separation and to estimate the size of molecular domains. The maximum path of diffusion in a 3D matrix, L , can be approximated via Equation 2.

$$L \sim \sqrt{6DT_1} \quad (2)$$

where D is the spin diffusion coefficient (for polymers, its value is approximately 5×10⁻¹⁶ m²·s⁻¹), which depends on the average distance between protons and dipolar interactions^[5,25].

2.5.3 Activation energy through spin-lattice relaxation

The activation energy values (E_a) for molecular relaxation processes occurring in the samples were determined through the method employed by^[27]. The spin-lattice relaxation rate can be described by Equation 3.

$$\frac{1}{T_1} = C \left[\left(\frac{\tau}{1 + \omega^2 \tau^2} \right) + \left(\frac{4\tau}{1 + 4\omega^2 \tau^2} \right) \right] \quad (3)$$

Table 2. Parameters employed in the inversion-recovery pulse sequence.

Parameters	Values
Pulse sequence	180°- τ -90°-Acquisition
90° pulse- μ s	Automatically set. Typically 7.5
180° pulse - μ s	Automatically set. Typically 15.
Number of scans	8
Number of τ (logarithmically spaced)	40
Tau sweep - μ s	10-2E7
Recycle delay - s	5
Receptor gain - %	3
Probe dead time - μ s	8
Receiver dead time – μ s	3

where τ is the correlation time; ω is the Larmor frequency; and C is a constant that represents the magnitude of the dipolar interaction between protons.

The correlation time can be understood as being thermally excited, as described in Equation 4.

$$\tau(T) = \tau_0 e^{\left(\frac{E_a}{RT}\right)} \quad (4)$$

where E_a is the activation energy; R is the ideal gas constant (here $8,314 \text{ J}\cdot\text{mol}^{-1}\cdot\text{K}^{-1}$); and τ_0 is a pre-exponential factor. The value of E_a can be obtained by plotting the natural logarithm of the relaxation constant versus the inverse of the absolute temperature^[27,28]. The linear regressions were conducted to obtain the coefficients of determination using a routine available in the Origin 8.5 software.

3. Results and Discussion

3.1 Sample analysis

Transparent, colorless, free-standing polymeric films were obtained for PVA and PVA/SiO₂ samples. Transparent, slightly yellow, free-standing films were obtained for all xerogels. The cross-linking reaction occurred when hydroxyl groups in the PVA and aldehyde groups in the GA underwent an acid-catalyzed acetalization process, producing acetal and hemiacetal bridges between polymer chains. This process is stochastic in nature and has already been shown to depend on factors such as polymer concentration and reactive group tacticity^[15,29,30]. A reaction scheme is shown in Figure 1.

Nanoparticulate hydrophilic SiO₂ is an amorphous material with many polar oxygenated groups on its surface. These groups promote high affinity for water, polar solvents and media, such as PVA. However, the SiO₂ used in this work has a very small surface area ($167.6 \text{ m}^2\cdot\text{g}^{-1}$, determined through BET N₂ sorption), so the extension of the interactions between nanoparticles and polymer matrix is expected to be small.

3.2 FT-IR analysis

FT-IR spectra of the samples can be observed in Figure 2a, b. PVA absorptions are reported as: secondary alcohol C-O stretching between $1124\text{--}1087 \text{ cm}^{-1}$; aliphatic C-H stretching between $2800\text{--}3000 \text{ cm}^{-1}$; and OH broad absorption around 3300 cm^{-1} , which contains contributions from hydrogen bonded moieties. The discrete absorption band at 1141 cm^{-1} was reported as being from stretching modes of C-C or C-O bonds, when PVA presents intramolecular hydrogen bonds^[31]. Therefore, this absorption band is associated with spatial order inside the polymer network. The absence of absorptions between $1750\text{--}1735 \text{ cm}^{-1}$, typical for ester C=O bands, confirms that this PVA is highly hydrolyzed^[31,32]. Comparing the spectra of PVA and of PVA/SiO₂ there are no apparent differences, which implies weak intermolecular interactions between the polymer matrix and inorganic nanoparticles (Figure 2a). For better visualization, we selected and expanded regions of the FT-IR spectra from PVA and P10 samples, as shown in Figure 3a, b. All the xerogels display absorption bands at $2850\text{--}2750 \text{ cm}^{-1}$ (Figure 3a) and $1720\text{--}1750 \text{ cm}^{-1}$ (Figure 3b), corresponding to absorption modes of short-chain aliphatic

aldehyde C-H and aldehyde carbonyls, which indicates some incomplete cross-linking. Although faint, the increase in the intensity of the absorption bands arising from cross-linking and the decrease in the intensity of the 1141 cm^{-1} band (Figure 3b), relative to spatial order in the PVA matrix, can be distinguished in the spectra. Quantification of the influence of these covalent bonds on the polymeric matrix proved to be somewhat difficult and prone to experimental errors. Furthermore, the technique did not show the apparent effect of the inorganic nanoparticles, due to the absence of strong chemical interactions.

3.3 WAXD analysis

Diffractograms of all the samples are displayed in Figure 4. Even though PVA does not form crystallites in the strict sense, it does display highly ordered domains that arise

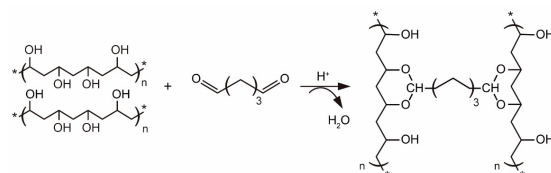


Figure 1. Reaction scheme for the formation of chemical cross-linking between PVA chains and GA.

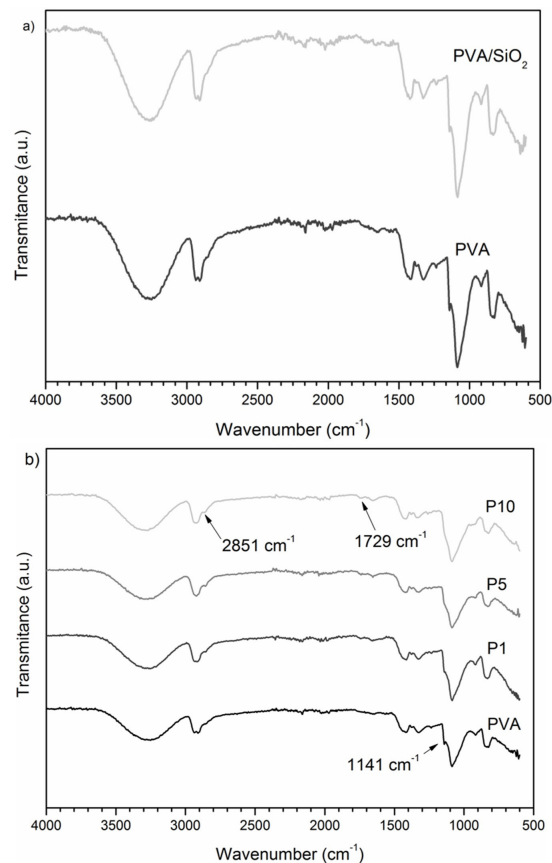


Figure 2. FT-IR spectra for the (a) non-cross-linked materials and (b) cross-linked materials.

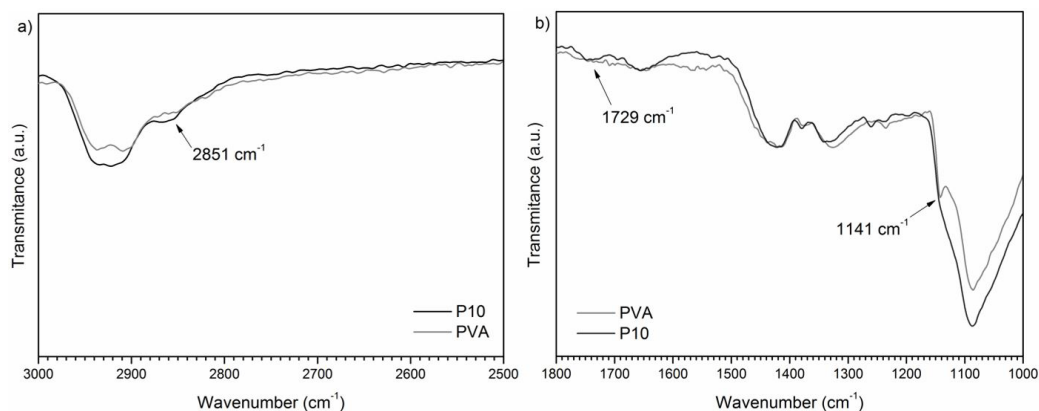


Figure 3. FT-IR spectra expansion of (a) short-chain aliphatic aldehyde C-H region and (b) aldehyde carbonyls and of C-C or C-O bonds.

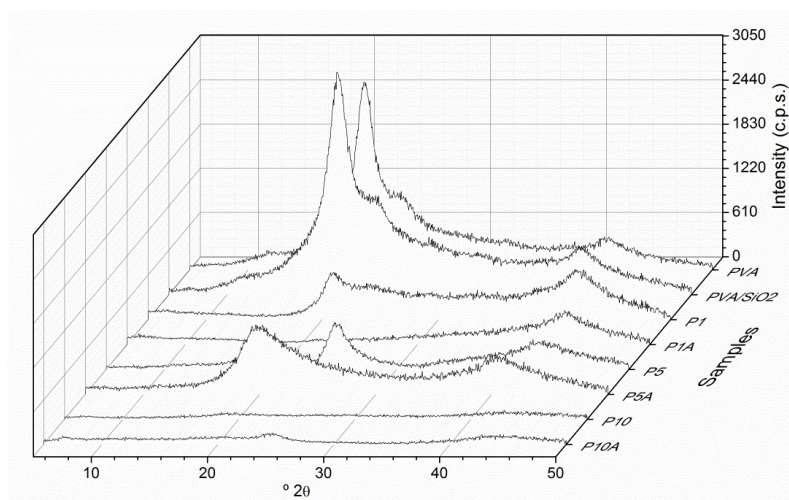


Figure 4. Superposed WAXD diffractograms of the samples.

because of the small size of the hydroxyl group, allowing inter/intramolecular hydrogen bonding and leading to more efficient packing of the polymer chains^[30,34]. The cross-linking reaction conducted in this work acts by eliminating free hydroxyl groups from the PVA chains, replacing them with acetal and hemiacetal bonds. This effect can be perceived when the xerogels' diffractograms are analyzed. The change is much more pronounced in the highly cross-linked samples. The presence of SiO₂ nanoparticles appears to have no effect on this change. At 10% nominal cross-linking, the intensity of the diffractions coming from the sample is so low that on the scale of the graph it cannot be easily distinguished from the baseline. It is clear that the ordered conformation of the PVA chains was undone by the cross-linking reaction, as this spatial order relies on the presence of hydroxyl groups. Thus, for these systems, WAXD measurements corroborated the results obtained through FT-IR, with respect to the strength of the cross-linking reaction.

3.4 TD-NMR analysis

All the acquisition parameters shown in Table 2 underwent various steps of iterative optimization to ensure better results and finer resolution. Once these parameters

are set, they can be repeated for each sample in a batch. When compared with the other characterization techniques, relaxometry offers some advantages. First, by observing only the spin-lattice relaxation processes we could extract up to three different and complementary readings about the samples. Second, TD-NMR probes the whole volume of the sample, providing more accurate and descriptive data about the material being analyzed. Third, if needed, the samples could have been used for other tests, since relaxometry is a nondestructive analytical tool^[1,3,4,7].

3.4.1 Spin-lattice relaxation constant determination

The spin-lattice relaxation constant, T_1 , can be interpreted as an estimate of the intermolecular interactions existing in the sample, since this parameter provides information about long chain segments under the frequency used in this experiment. This time constant is very dependent on the correlation times of the proton nuclei in the PVA chains, so that nuclei constrained in rigid domains display long correlation times. This, in turn, translates into greater values of T_1 . The inverse holds for short T_1 values, which arise from mobile, less restricted molecular environments^[18,19,33].

The T_1 values for PVA and PVA/SiO₂ were used as standards against which the values for xerogels were measured. Tables 3 and 4 show the T_1 values for the samples. The data are grouped in filled and non-filled samples, for clarity. The magnetization relaxation back to thermal equilibrium via the T_1 mechanism is modulated by the molecular mobility (more or less restrained polymer chains) of the sample. The value of T_1 decreases with increasing temperature, suggesting a reduction of molecular rigidity of the PVA matrix, which is expected since the system receives thermal energy. This trend occurs despite the degree of cross-linking or the presence of inorganic nanoparticles. The presence of nanoparticles, however, seems to have a dampening effect on the reduction of the samples' T_1 values due to this stimulus, especially at high temperatures. Another trend is the decrease of the T_1 values with increased cross-linking between PVA chains, although the intensity of the change in T_1 values starts to diminish when the cross-linking degree reaches 5% (m/m), suggesting saturation around this concentration of GA. This behavior is similar to that described by Zhao et al.^[18]. The main difference between our results and theirs is that we investigated a vinyl polymer and they examined an elastomer. As previously discussed, the cross-linking process occurs, for the systems studied here, by consumption of hydroxyl moieties present in the polymer. According to the experimental results, we attributed greater lattice rigidity to neat PVA due to the strong hydrogen bonding from hydroxyl moieties in its structure. The conversion of said moieties to (hemi)acetal groups linked by three CH₂ bonds, which are somewhat mobile, thus increased the net molecular mobility of the polymer matrix. That being said, the presence acetal bridges diminishes spatial organization and molecular rigidity, as reflected in the T_1 values. Furthermore, the nanoparticles' influence could be perceived when looking at T_1 values for a given neat polymer/hybrid pair. For all the formulations, the hybrid displayed a shorter T_1 value than its unloaded equivalent, thus showing that T_1 is sensitive to the very weak surface interactions between the polymer matrix and the SiO₂ nanoparticles. This result is interesting because it allows faster and better interpretation of the extent of the changes brought by the addition of the nanoparticles than the other two common spectroscopic methods employed.

3.4.2 Spin diffusion path analysis

Regarding homogeneity at the molecular level, spin-lattice relaxation can be used to describe the extent of similar spin environments at nanometer scales. The parameter L is the theoretical distance where spin diffusion events can take place without causing the averaging of the magnetization. Thus, this value represents the upper limit of similar molecular mobility. In other words, L is an approximation of the size of homogeneous domains inside the material^[5,25].

The values for L shown in Figure 5 demonstrate a downward trend with rising quantity of nanoparticles and cross-linking degree. The effect of temperature on this parameter is also demonstrated, with L values for the samples evaluated at 35 and 55 °C.

As the results show, the introduction of acetal bridges breaks the spatial organization between PVA chains, and although they are stochastic in nature, the effects of the cross-linking reaction demonstrate an almost linear decay with increasing GA concentration. Comparing PVA with PVA/SiO₂, the introduction of SiO₂ does little to disturb the homogeneity of PVA chains. Even the smallest amount of GA (P1) had a more pronounced effect on this parameter than that caused by the nanoparticles' presence. Finally, the technique is sensitive to the effects of temperature on the homogeneity of the polymeric spin domains. With the increase of thermal energy, random movements of long

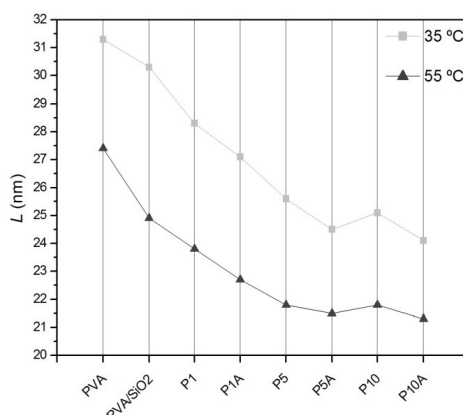


Figure 5. Spin diffusion path for all samples.

Table 3. T_1 values for the non-filled samples at different temperatures.

Sample	T_1 (ms) at 35 °C	T_1 (ms) at 45 °C	T_1 (ms) at 55 °C	T_1 (ms) at 65 °C
PVA	326	286	251	188
P1	267	226	189	142
P5	219	188	159	121
P10	211	188	159	120

Table 4. T_1 values for the filled samples at different temperatures.

Sample	T_1 (ms) at 35 °C	T_1 (ms) at 45 °C	T_1 (ms) at 55 °C	T_1 (ms) at 65 °C
PVA/SiO ₂	306	243	207	197
P1A	246	205	172	127
P5A	200	176	154	120
P10A	193	172	152	116

Table 5. E_a values calculated for all the samples.

Sample	E_a (kJ.mol ⁻¹)	R ²
PVA	15.36	0.99
PVA/SiO ₂	12.93	0.98
P1	14.48	0.99
P1A	15.02	0.99
P5	13.35	0.99
P5A	10.93	0.99
P10	11.81	0.97
P10A	10.05	0.99

segments of polymer chains contribute to reduction of the spin diffusion path's length^[5,25,26].

3.4.3 Activation energy through spin-lattice relaxation

Values for E_a are related to the polymer matrix lateral groups' rotation movement and to chain end or ramification vibrations^[27,28,33]. We treated this quantity as apparent energy of activation. Table 5 lists the values for all samples.

Both cross-linking and the presence of nanoparticles caused a reduction in the value of the systems' apparent activation energy, albeit through different mechanisms. As seen before, the SiO₂ particles do not interact chemically with the polymeric matrix, but their presence is sufficient to reduce the size of the homogeneous spin domains. The reduction on the E_a value for PVA/SiO₂ must have come from a disorganized, somewhat less restricted environment. The same rationale explains the effects of cross-linking, as the reaction consumes hydroxyl moieties, reported to be responsible for spatial ordering of PVA chains^[30,31]. The results obtained for the apparent E_a values of non-filled samples indicate these values can be used to gauge the extent of cross-linking in this type of xerogel, as they vary linearly with GA concentration.

4. Conclusions

We prepared and analyzed SiO₂ loaded and neat PVA based xerogels through FT-IR, WAXD and TD-NMR, by means of spin-lattice relaxation mechanisms. The results obtained from the FT-IR and the WAXD analyses demonstrated that the cross-linking reaction reduces the spatial ordering of the PVA polymer chains. There was no chemical interaction between the PVA hydroxyl groups and the SiO₂ surface that could be detected via FT-IR, although TD-NMR results did show some level of physical interaction between polymer and nanoparticles. All three TD-NMR experiments performed in this study revealed the effect of the SiO₂ presence amid the PVA chains. The data derived from the TD-NMR experiment, such as E_a and T_1 constant values, can be readily employed to assess this kind of material's degree of cross-linking without resorting to swelling-to-equilibrium solvent intake tests, with the added advantage of being faster. The experimental design employed in the work enables clearer and broader characterization of an important class of polymeric materials, the inorganic-organic hybrid xerogels.

5. Acknowledgements

We thank Roberto Pinto Cucinelli Neto for the fruitful discussions about relaxometry and his help during some of the analyses. We also thank Coordenação de Aperfeiçoamento de Pessoal de Ensino Superior (CAPES), Conselho Nacional de Desenvolvimento Científico e Tecnológico (CNPq) and Fundação Carlos Chagas Filho de Amparo a Pesquisa do Estado do Rio de Janeiro (FAPERJ) for funding.

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Received: Mar. 06, 2015

Revised: Oct. 21, 2015

Accepted: Feb. 15, 2016