

## IMPROVING THE SOLUBILITY OF THE ANTICHAGASIC DRUG BENZNIDAZOLE THROUGH FORMATION OF INCLUSION COMPLEXES WITH CYCLODEXTRINS

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This study describes unpublished research on improving the solubility of benznidazole by the formation of an inclusion complex. The cyclodextrins selected were  $\alpha$ CD,  $\beta$ CD,  $\gamma$ CD, HP $\beta$ CD, RM $\beta$ CD and SB $\beta$ CD. All complexes were obtained in solution, presenting 1:1 stoichiometry according to the phase solubility diagram. The highest association constants were obtained with RM $\beta$ CD and SB $\beta$ CD, being selected for attainment of solid state complexes. These were characterized using XRD, SEM and dissolution test. The data obtained suggest the formation of complexes and indicate that these may provide a promising alternative way of developing solid doses of drug with suitable biopharmaceutical properties.

Keywords: benznidazole; Chagas disease; inclusion complex and dissolution.

### INTRODUCTION

Chagas disease is an endemic disease caused by the protozoan *Trypanosoma cruzi* and is considered to be a serious medical and social problem in Latin America, affecting approximately 10 million individuals, with an additional 40 million at risk of infection.<sup>1,2</sup>

Although it is not the ideal drug of choice, owing to its low solubility and consequent high toxicity, benznidazole (BNZ), N-benzyl-2-(2-nitroimidazol-1-yl) acetamide, is currently the only drug available for Chagas disease treatment.<sup>3</sup>

Increased knowledge of the physicochemical characteristics of drugs and of more effective techniques for improving the biopharmaceutical properties of drugs enables new therapeutic alternatives to be developed more cheaply and more quickly than the introduction of a new molecule into the treatment.

Several techniques, such as solid dispersion,<sup>4</sup> sol-emulsion,<sup>5</sup> inclusion complexation and conjugation with water-soluble polymer or cyclodextrin have been developed to enhance the solubility of various drugs.<sup>6-11</sup>

The potential use of natural cyclodextrins (CDs) and their synthetic derivatives has been extensively studied as a way of improving certain drug properties, such as solubility, stability and/or bioavailability. The property of enhancing solubility may be explained by the formation of water-soluble inclusion complexes in which the hollow, truncated, cone-like CD structure encapsulates the hydrophobic drug molecules in the apolar interior. By contrast, the outer, hydrophilic region of the CD, enables solubilization through interaction with water molecules.<sup>12,13</sup>

As yet few researchers have adopted technical approaches that aim to optimize the biopharmaceutical properties of BNZ and none of these

have considered the development of a solid pharmaceutical form. Alternatives to a parenteral solution of BNZ found in the literature include the development of liposomes<sup>2,4</sup> and obtention of a ruthenium complex, *trans*-[Ru(Bz)(NH<sub>3</sub>)<sub>4</sub>SO<sub>2</sub>](CF<sub>3</sub>SO<sub>3</sub>)<sub>2</sub><sup>14</sup> which alters the solubility of the compound and increases the effectiveness of the drug. However, no study involving BNZ has placed the drug in an inclusion complex, an approach that may hold promise for the development of a solid dosage form that is more appropriate than the one currently available.<sup>15</sup>

The purpose of this study was to improve the solubility and dissolution rate of BNZ in order to develop a promising solid oral delivery system. Firstly, a screening of six types of CDs was performed and their capacity of increasing the drug intrinsic solubility evaluated by the phase solubility studies, being three natural CDs ( $\alpha$ ,  $\beta$  and  $\gamma$ ) and three modified  $\beta$ CDs [randomly methylated  $\beta$ CD (RM $\beta$ CD), hydroxypropyl  $\beta$ CD (HP $\beta$ CD) and sulfobutyl ether  $\beta$ CD (SB $\beta$ CD)]. In a second phase, BNZ-CD binary systems were prepared in solid state. Only the two modified  $\beta$ CDs, RM $\beta$ CD and SB $\beta$ CD, which had the highest complex association constants for this phase. The solid state complexes were obtained using a variety of methods, including physical mixing, kneading, freeze-drying and evaporation. The inclusion complexes obtained were characterized using X-ray powder diffraction (XRPD) and scanning electron microscopy (SEM), and the dissolution profile of the complex and the drug were determined.

### EXPERIMENTAL

#### Materials

The BNZ, N-benzyl-2-(2-nitroimidazol-1-yl) acetamide, was supplied by Roche. The purity was estimated using DSC and high performance liquid chromatography (HPLC) to be 99.9%.  $\beta$ CD and RM $\beta$ CD (degree of molar substitution of 0.57) were donated by

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Roquette (Barcelona, Spain); HP $\beta$ CD (degree of molar substitution of 2.7) by Janssen Pharmaceutica (Beerse, Belgium) and SB $\beta$ CD (degree of molar substitution of 1.0) by CyDex (Lenexa, KS). All other materials were of analytical grade. The solutions were prepared using distilled water, filtered through 0.22  $\mu$ m Millipore filters (Millipore Corp<sup>®</sup>, Billerica, MA).

### Analytical methods

The UV spectrophotometric method was used, following the method previously reported by Soares-Sobrinho *et al.*<sup>16</sup> for quantitative determination of BNZ, on an Agilent 8453 UV-Visible spectrophotometer (Agilent Corp, Santa Clara, CA) with a photodiode array detector set at 324 nm. The calibration curve in water was plotted using standard solutions in the range of 4 to 40  $\mu$ g mL<sup>-1</sup>. The addition of CD's was observed to have no effect on the UV spectrum of the BNZ solution for any of the CDs studied.

### Phase solubility diagram

Phase solubility studies were performed according to the method reported by Higuchi *et al.*<sup>17</sup> Excess quantities of BNZ (~30 mg) were added to 5 mL aqueous solutions containing increasing concentrations of CDs, in the range of 0 to 1.5% (wt/vol) for  $\beta$ CD; 0 to 2% (wt/vol) for  $\alpha$  and  $\gamma$  CD; 0 to 4% (wt/vol) for HP $\beta$ CD RM $\beta$ CD and SB $\beta$ CD. Suspensions were placed in an ultrasound bath for 15 min and shaken in an oscillating water bath thermostatically controlled at 25 °C for 7 days, which previous studies have shown to be sufficient time to ensure equilibrium. Samples were filtered and appropriately diluted. The concentration of BNZ was determined using a spectrophotometer. Experiments were performed in triplicate. The drug-CD association constants ( $K_{1:1}$ ) were calculated using the linear portion of the phase solubility diagrams, assuming 1:1 stoichiometry, according to the Equation 1, where  $S_0$  is BNZ water solubility in the absence of CD:

$$K_{1:1} = \frac{\text{Slope}}{S_0 \times (1 - \text{Slope})} \quad (1)$$

### Preparation of solid-state complexes

#### Preparation of physical mixture

A physical mixture (PM) of BNZ and the CDs RM $\beta$ CD or SBE $\beta$ CD was obtained by mixing the compounds in a Turbula WAB T2C mixer (Willy A. Bachofen Corp, Basel, Switzerland) at a 1:1 mol ratio for 25 min.

#### Preparation of kneaded systems (KN)

The kneaded system (KN) was prepared from the physical mixture in a mortar by the slow addition of ethanol/water (1:1 vol/vol), at 20% wt/wt, and mixed until a homogeneous system was obtained. Samples were dried for 24 h in a 40 °C oven and powdered.

#### Preparation of freeze-dried systems (FD)

The physical mixture was dissolved in ethanol:water (1:1 vol/vol). The solution was frozen by immersion in liquid nitrogen and freeze-dried (FD) in a Labconco apparatus (Labconco Corp, Kansas City, MO).

#### Evaporation (EV)

The physical mixture was dissolved in ethanol:water (1:1 vol/vol). The solution was evaporated in a vacuum at 50 °C in a rotary evaporator (Büchi Rotavapor R-200).

### The characterization of the solid-state complexes

#### X-ray powder diffraction (XRPD)

The X-ray powder diffraction patterns were collected using copper radiation (40 Kv, 30 mA), on a Philips PW 1729 diffractometer (Philips Corp, Netherlands) with Bragg-Brentano geometry, in the 2<math>2\theta</math><math>60</math> range with a step size of 0.02° and counting time of 2 s per step.

#### Scanning electron microscopy (SEM)

The surface morphologies of samples were examined using a LEO-435VP scanning electron microscopy (Leica Microsystems, Cambridge, UK). Particles were fixed on a brass stub using double-sided tape and vacuum coated gold.

#### Dissolution studies

*In vitro* dissolution studies were performed in accordance with the Food and Drug Administration's<sup>18</sup> specifications for poorly soluble drugs. The United States Pharmacopeia<sup>19</sup> dissolution two paddle apparatus (Turu Grau DT-6, Barcelona, Spain) set at 75 rpm and 37 °C ( $\pm$  0.5 °C) was used. In order to ensure the *Sink condition* of the method and increase the specificity of the dissolution test, to show that this method is able to distinguish between different drug formulations, the recommended medium for poorly soluble drugs was selected: namely, sodium chloride : hydrochloride acid buffer pH 1.2, with a volume of 900 mL. Powdered samples of 50 mg of BNZ or equivalent amounts of each BNZ/CD system (particle size 50-200  $\mu$ m) were tested. Samples were collected at certain time intervals, filtered through cellulose filters, and the concentration of dissolved drug was determined by spectroscopy at 324 nm. Three replicates were tested for each binary system. The dissolution profiles were assessed and compared using dissolution efficiency at 15 and 20 min (DE15 and DE20) as the parameters.<sup>20</sup> The statistical analysis of the DE20 was performed using one-way analysis of variance (ANOVA) followed by a least significant difference (LSD) test using Statgraphics Plus Version 3.

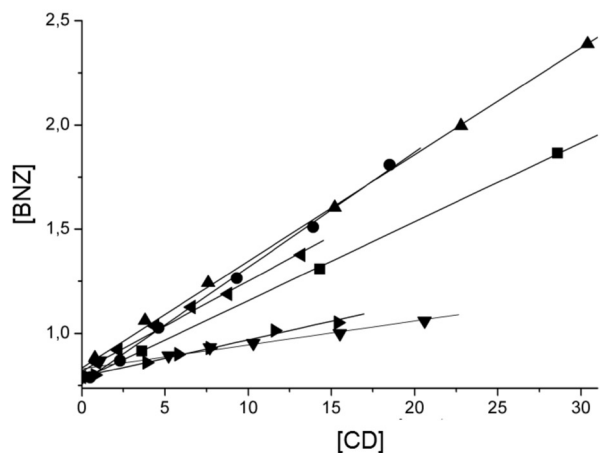
## RESULTS AND DISCUSSION

### Phase solubility diagram

In this paper, we have used a method to determinate the association constants of the inclusion complexes formed between BNZ and all CDs used, based on phase solubility studies (Figure 1). The improvement in solubility of the BNZ depended greatly on the type of CD. In all cases, the solubility of BNZ increased linearly as a function of CD concentration over the concentration range under study, with gradients lower than one. These diagrams can be classified as  $A_L$  type and suggest a 1:1 mol:mol BNZ-CD stoichiometry, according to Higuchi *et al.*<sup>17</sup> This drug had not hitherto been evaluated using studies involving CDs.

The  $K_{1:1}$  values obtained for  $\beta$ CD and derivatives (HP $\beta$ CD, RM $\beta$ CD, SB $\beta$ CD) were almost twice as high as those obtained for the other CDs ( $\alpha$ CD and  $\gamma$ CD) (Figure 1) (Table 1). The improved solubility of RM $\beta$ CD may result from the presence of methyl groups, which can extend the hydrophobic region of the CD cavity, facilitating and stabilizing BNZ inclusion complexation.<sup>21</sup> A similar effect also occurs with the hydroxypropyl and sulfobutyl ether groups. All complex solutions had a final pH of  $\approx$  7.0, indicating no change in pH occurred during the study.

As mentioned above, for a solid oral formulation it would be useful to select the CD derivatives (RM $\beta$ CD, SB $\beta$ CD) with the greatest capacity for increasing BNZ solubility. Therefore, the general



**Figure 1.** Phase solubility diagrams of BNZ as a function of cyclodextrin concentrations at 25 °C in water. (■) HPβCD, (●) SBβCD, (▲) RMβCD, (▼) αCD, (▴) γCD, (◄) βCD

**Table 1.** Results obtained from the phase solubility diagram

CDs	$K_{1:1}$ (mol <sup>-1</sup> )	R <sup>2</sup>
HPβCD	49.7	0.9992
SBβCD	74.1	0.9985
RMβCD	68.3	0.9982
αCD	14.7	0.9699
γCD	22.9	0.9579
βCD	57.2	0.9981

characteristics of RMβCD and SBβCD, especially their high aqueous solubility and low toxicity, in addition to their greater capacity to dissolve BNZ, made these CDs the candidates for inclusion in an oral solid dosage form.<sup>22</sup>

### Characterization of solid-state complexes

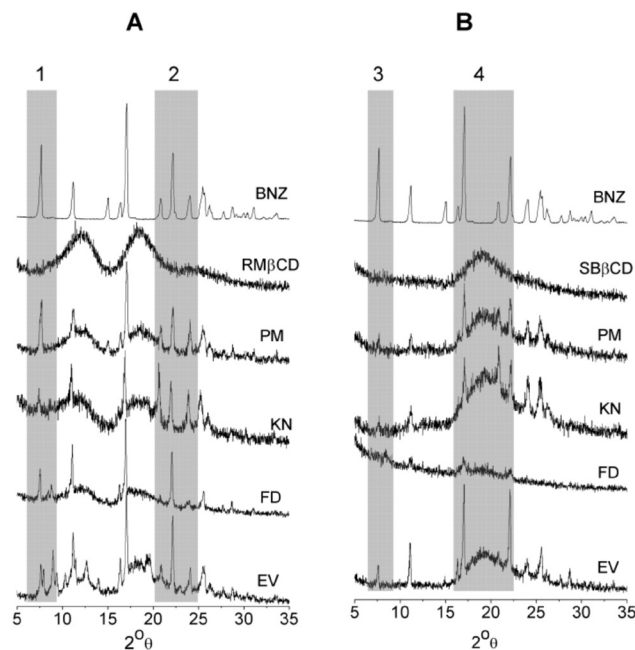
#### X-ray powder diffraction

XRPD was carried out on the samples in isolation (BNZ, RMβCD and SBβCD), PM and EV, FD and KN, with a view to identifying evidence of the formation of CD inclusion complexes (Figure 2). The diffractometric profile of the drug reveals the presence of high-intensity peaks at 7.5, 16.5 and 22°2θ, along with a number of secondary peaks, thereby demonstrating the crystalline behavior of BNZ. The opposite can be seen in the diffratograms of the predominantly amorphous CDs.

The PM method of preparation, both for RMβCD and SBβCD, correspond to the sum of both profiles with a reduction in the intensity of BNZ peaks owing to its concentration in the sample. It is possible, according to Figure 2, to identify alterations in some characteristic peaks of BNZ in addition to the disappearance or lower intensity of the other peaks in the processed preparations (KN, FD and EV). This may be taken as evidence of interaction between the components being facilitated by the processing of samples.

Column 1 in Figure 2A shows the lower peak in the KN, FD and EV preparations, when compared to PM, along with the appearance of new peaks near the characteristic initial peak of BNZ at 7.5°2θ. The appearance of new peaks not corresponding to the original component provides strong evidence of the formation of a new crystalline phase that may correspond to the drug-CD inclusion complex.

Column 2 in Figure 2A points to the disappearance and weakening of peaks in the region between 20 and 26°2θ for the FD sample, which



**Figure 2.** X-ray diffraction patterns of BNZ with RMβCD (A) and SBβCD (B) as supplied and their processed binary systems

may indicate loss of crystalline structure in this sample, which can also be interpreted as evidence of interaction between BNZ and RMβCD.

It was possible to detect small variations in the relative intensity of the peaks for the drug in mixtures containing SBβCD (Column 3-4, Figure 2B) in the processed samples. However, these variations are difficult to interpret, since they may be caused by an effect of the preferred orientation of crystals. For the FD preparation (Figure 2B), a more pronounced amorphous profile can be seen, with alterations to the characteristics of the main crystalline peaks of the drug and, in the case of the preparation illustrated in Figure 2A, this may indicate interaction between components.

In view of these changes observed from the diffratograms, it is fair to say that there is consistent evidence of interaction and, consequently, possible formation of inclusion complexes.<sup>23,24</sup>

#### Scanning electron microscopy (SEM)

Morphological changes in particles can have an important impact on the dissolution profile.<sup>25,26</sup> Selected photomicrographs of BNZ, RMβCD and SBβCD, along with those of binary systems processed are represented in Figure 3. BNZ characteristically takes the form of regular crystals,<sup>27</sup> while RMβCD and SBβCD are normally composed of amorphous particles.<sup>8</sup>

In general, all the processed samples underwent changes in the morphology of the particles. The characteristics of the amorphous particles of the non-manipulated CDs were not observed. The KN preparation form aggregates (granules) that are somewhat uniform in size. The FD formulation of both CDs and the EV formulation of RMβCD show a reduction in the size of particles, which take the form of needles shaped completely differently from the materials at the start of the experiment. The change in shape and reduction in size of particles seems to be less striking for the EV preparation of SBβCD.

#### Dissolution studies

The dissolution profiles for the systems under study are presented in Figure 4. The dissolution efficiency was calculated after 15 and 20 min for all the products studied (Table 2). Data from the literature show that increases in the dissolution of inclusion complexes for

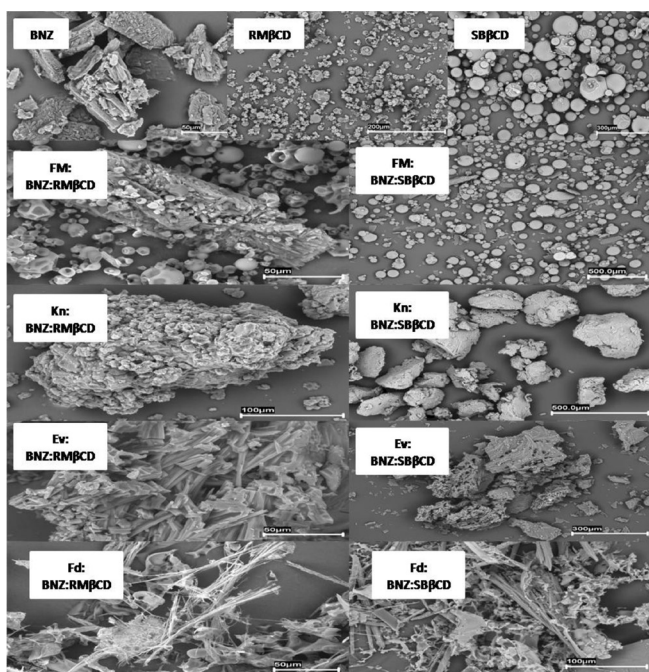


Figure 3. Scanning electron microphotographs of BNZ; RM $\beta$ CD; SB $\beta$ CD; PM; KN, EV and FD systems

different drugs can be explained by their higher energy amorphous state (reduction in crystallinity) and by the lower interfacial tension between water and the insoluble drug.<sup>25</sup>

Significant differences were observed in the dissolved quantities of pure drug (BNZ), PM and processed forms (KN, FD and EV) as well as in the dissolution rate. Five minutes after the start of the dissolution test, the percentage of drug dissolved was  $\approx 12\%$ ; the percentage of drug dissolved from the RM $\beta$ CD complex was over 80% and the percentage of drug dissolved from the SB $\beta$ CD complex was over 50%.

In the case of both PMs, the dissolution efficiency brought about a significant difference in comparison to pure BNZ ( $p < 0.05$ ). This appears to be a consequence of the mechanical treatment which increases the contact between the drug and the carrier. This fact can be attributed to the improved wettability of PMs, owing to the presence of cyclodextrins,<sup>28</sup> as well as, in the earlier stages of the dissolution process, CD molecules operating locally on the hydrodynamic layer surrounding the particles of BNZ. This results in an *in situ* inclusion process and produces a rapid increase in the amount of dissolved drug.<sup>29</sup>

Although the kneading process is efficient in improving the dissolution rate for binary systems, as in the case of the BNZ-RM $\beta$ CD complex, where  $\approx 100\%$  of the drug dissolved within 10 min of testing, the interaction with the drug requires a better method, such as the FD or EV, to obtain similar results for mixtures with SB $\beta$ CD. These results may be explained by an intrinsic interaction between BNZ and RM $\beta$ CD at surface level, as demonstrated by SEM (Figure 4), as this facilitates the formation of the complex.

The dissolution rates ranked as follows according to ANOVA: BNZ < PMSB $\beta$ CD < PMRM $\beta$ CD = KNSB $\beta$ CD < KNRM $\beta$ CD = FD = EV. These results suggest FD and EV should be the recommended

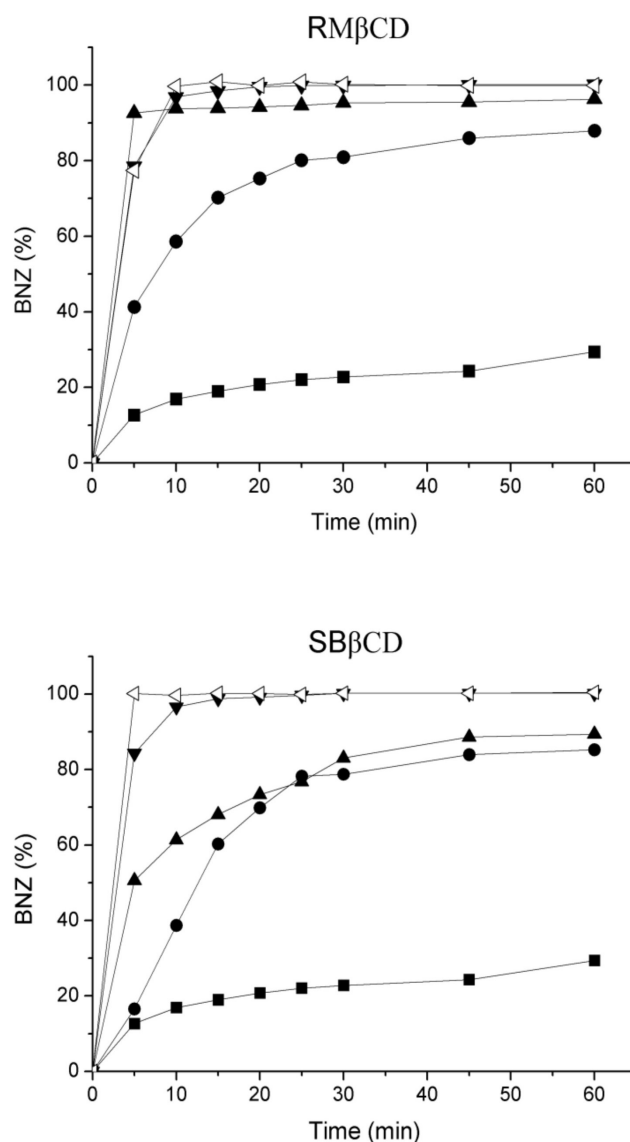


Figure 4. Dissolution profiles of BNZ, PM, KN, EV and FD systems with RM $\beta$ CD and SB $\beta$ CD, in sink condition. (■) BNZ, (●) PM, (▲) KN, (▼) FD, (◄) EV

methods for preparation of inclusion complexes between BNZ and CDs with a higher solubilizing power. The KN method could also be used for BNZ:RM $\beta$ CD binary mixtures, owing to the high capacity of CD to form inclusion complexes spontaneously.

## CONCLUSION

In conclusion, the  $\beta$ CD derivatives, SB $\beta$ CD and RM $\beta$ CD, can be shown to be highly efficient in terms of BNZ solubilization. The results suggest the possible formation of inclusion complexes using the EV, FD and KN techniques, with a dramatic improvement in drug dissolution rate compared to the drug alone or the CD/drug PM. The SEM and XRPD results obtained show that a BNZ:CDs inclusion

Table 2. Calculated mean values of dissolution efficiency at 15 and 20 min

	BNZ	PM		KN		FD		EV	
		RM $\beta$ CD	SB $\beta$ CD	RM $\beta$ CD	SB $\beta$ CD	RM $\beta$ CD	SB $\beta$ CD	RM $\beta$ CD	SB $\beta$ CD
%DE <sub>15min</sub>	12.4 ( $\pm 1.2$ )	46.9 ( $\pm 2.6$ )	28.8 ( $\pm 3.7$ )	77.8 ( $\pm 1.2$ )	48.7 ( $\pm 2.1$ )	74.7 ( $\pm 2.8$ )	76.1 ( $\pm 2.0$ )	83.4 ( $\pm 1.3$ )	75.8 ( $\pm 1.3$ )
%DE <sub>20min</sub>	14.2 ( $\pm 1.3$ )	54.4 ( $\pm 1.9$ )	37.8 ( $\pm 3.7$ )	81.8 ( $\pm 1.1$ )	54.2 ( $\pm 1.9$ )	81.0 ( $\pm 1.7$ )	81.8 ( $\pm 2.2$ )	87.6 ( $\pm 1.8$ )	81.9 ( $\pm 1.5$ )



complex with a good dissolution profile could be prepared at a 1:1 molar ratio using EV, FD and KN methods. In view of these results, we believe that complexation effectively enhanced the solubility of BNZ, which may consequently increase its bioavailability and might improve its pharmaceutical potential, providing a more uniform and efficient dissolution profile, minimizing variations during the drug solubilization stage, exposing the organism to lower quantities of the drug and hence lowering its toxicity.

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