

# Inclusion complexes and self-assembled cyclodextrin aggregates for increasing the solubility of benzimidazoles

Leticia Norma Carpentieri Rodrigues<sup>1\*</sup>, Anna Carolina Moraes Tavares<sup>1</sup>, Beatriz Tavares Ferreira<sup>1</sup>, Adriana Karla Cardoso Amorim Reis<sup>1</sup>, Luciana Morita Katiki<sup>2</sup>

<sup>1</sup>Instituto de Ciências Ambientais, Químicas e Farmacêuticas, Universidade Federal de São Paulo, Diadema, São Paulo, Brazil, <sup>2</sup>Secretaria de Agricultura e Abastecimento do Estado de São Paulo, Instituto de Zootecnia – IZ, Agência Paulista de Tecnologia dos Agronegócios – APTA, Nova Odessa, São Paulo, Brasil.

Albendazole and fenbendazole are imidazole derivatives that exhibit broad spectrum activity against parasites, but the low solubility of these drugs considerably reduces their effectiveness. Complexation of albendazole and fenbendazole with cyclodextrins ( $\beta$ -cyclodextrin and hydroxypropyl- $\beta$ -cyclodextrin) in both water and an aqueous solution of polyvinylpyrrolidone (PVP-k30) was studied to determine if it could increase the solubility and dissolution rate of the drugs. In an aqueous solution,  $\beta$ -cyclodextrin increased the solubility of albendazole from 0.4188 to  $\sim 93.47 \mu\text{g mL}^{-1}$  (223 $\times$ ), and of fenbendazole from 0.1054 to  $45.56 \mu\text{g mL}^{-1}$  (432 $\times$ ); hydroxypropyl- $\beta$ -cyclodextrin, on the other hand, increased solubility to  $\sim 443.06 \mu\text{g mL}^{-1}$  (1058 $\times$ ) for albendazole and  $\sim 159.36 \mu\text{g mL}^{-1}$  (1512 $\times$ ) for fenbendazole. The combination of hydroxypropyl- $\beta$ -cyclodextrin and polyvinylpyrrolidone enabled a solubility increase of 1412 $\times$  ( $\sim 591.22 \mu\text{g mL}^{-1}$ ) for albendazole and 1373 $\times$  ( $\sim 144.66 \mu\text{g mL}^{-1}$ ) for fenbendazole. The dissolution rate of the drugs was significantly increased in binary and ternary systems, with hydroxypropyl- $\beta$ -cyclodextrin proving to be more effective. The presence of the water-soluble PVP-k30 increased the dissolution rate and amorphization of the complexes. Analysis of the changes in displacement and the profile of the cyclodextrin bands in the <sup>1</sup>H NMR spectra revealed a molecular interaction and pointed to an effective complexation in the drug/cyclodextrin systems. Monomeric forms and nanoclusters of cyclodextrins were observed in the drug/cyclodextrin systems, suggesting that the increase in solubility of the drugs in the presence of cyclodextrins should not be attributed only to the formation of inclusion complexes, but also to the formation of cyclodextrin aggregates.

**Keywords:** Cyclodextrins/pharmacokinetics. Benzimidazoles/administration & dosage. Inclusion complexes. Cyclodextrin aggregates. Drug liberation/drug effects. Antiparasitic agents/analysis. Albendazole/analysis. Fenbendazole/analysis.

## INTRODUCTION

Benzimidazoles (BNZs) are a class of chemotherapeutic agents with broad biological activity. They are employed as antiparasitic agents and antibacterial agents, and recently, important antitumor activity has been described (Duan, Liu, Rockwell, 2013). In addition, nitroaromatic derivatives of benzimidazole are being studied for the treatment of tuberculosis and malaria (Camacho *et al.*, 2011). Despite its wide-ranging biological activities,

the low solubility of this class of substances has resulted in its inefficiency as a therapeutic agent, and parasitic and bacterial resistance against it has emerged.

Different strategies have been investigated to solve the limited aqueous solubility of benzimidazoles, including the use of surfactants, such as polysorbates and bile salts, or co-solvents (Daniel-Mwambete *et al.*, 2004; Miranda *et al.*, 2011; Pradines *et al.*, 2014). There is evidence that complexation with cyclodextrins (CDs) can increase the oral bioavailability of benzimidazoles (Evrard *et al.*, 2002; Moriwaki *et al.*, 2008; Pourgholami, Wangoo, Morris, 2008; Palomares-Alonso *et al.*, 2010; Kumar *et al.*, 2011a; Kumar *et al.*, 2011b; Alamdarnejad *et al.*, 2013; De Jaeghere *et al.*, 2013; Pradines *et al.*, 2014; Ferreira *et al.*, 2015). CDs are known for increasing

\*Correspondence: L. N. C. Rodrigues. Instituto de Ciências Ambientais, Químicas e Farmacêuticas, Universidade Federal de São Paulo. Rua São Nicolau, 210, 09913-030, Diadema, São Paulo, Brazil. Telephone/Extension line: +55 11 4044-0500/3576. E-mail: leticia.carpentieri@gmail.com

the solubility, dissolution efficiency, and bioavailability of different drugs through the formation of inclusion complexes. Recently, it has been observed that CDs are capable not only of forming inclusion complexes, but also of self-organizing to form nanoclusters that can contribute to the dissolution of low solubility drugs (Heydari, Kakhki, 2013; Kurkov, Loftsson, 2013).

Stella and Rajewski (1997) have pointed out that the mechanism involved in the increase of BNZ solubility by the addition of CDs has advantages compared to that by the addition of co-solvents and surfactants. In the case of co-solvents and surfactants, the increase in solubility occurs through changes in the properties of the solution, and any change in the medium can compromise the solubilization of the drug. On the other hand, dissolution through the formation of inclusion complexes with CDs is not influenced by the medium.

The objective of this study was to develop and characterize the inclusion complexes of benzimidazole-2-carbamates (albendazole and fenbendazole) and cyclodextrins, and to evaluate the effect of the cyclodextrins on the solubility of these drugs.

## MATERIALS AND METHODS

### Chemicals and reagents

Albendazole (ABZ), fenbendazole (FBZ), polyvinylpyrrolidone (PVP-k-30, molecular weight 360,000) and hydroxypropyl- $\beta$ -cyclodextrin (HP $\beta$ CD) were obtained from Sigma-Aldrich Co. (São Paulo, Brazil).  $\beta$ -Cyclodextrin ( $\beta$ CD, molecular weight 1135) was kindly provided by Labonathus Health Solutions (São Paulo, Brazil). Acetonitrile (HPLC grade) was obtained from J. T. Baker. Ultra-purified water was obtained through the Thermo Scientific reverse osmosis system. All other reagents were of analytical grade.

### HPLC determination

The criteria used were based on the methodology described by Rodrigues, Reis, and Katiki (2016).

The BNZs (ABZ or FBZ) were quantified using an HPLC (Shimadzu Prominence Modular HPLC) with a diode-array detector set at a wavelength of 300 nm, a mobile phase of acetonitrile:0.04 M diammonium hydrogen phosphate (41:59 v/v) with pH 7.5-7.6, a flow rate of 1.5 mL min<sup>-1</sup>, an injection volume of 20  $\mu$ L, and a reverse phase chromatographic column (Kromazil reverse RP C18, 100 A, 5  $\mu$ m, 250  $\times$  4.6 mm) equilibrated to ambient temperature and eluted under isocratic conditions.

The mobile phase was previously degassed through vacuum filtration (KNF vacuum pump N810.FT.18, 100 mbar) through a 45  $\mu$ m Nylon membrane, followed by sonication in an ultrasonic bath (Altasonic 9L) for 15 min.

Calibration curves were obtained using stock solutions of the drugs (ABZ or FBZ, 100  $\mu$ g mL<sup>-1</sup>). Linear relationships were obtained in the range of 2 to 100  $\mu$ g mL<sup>-1</sup> for both drugs; the limit of detection (LOD) and limit of quantification (LOQ) values obtained for ABZ were 0.4855 and 1.6183  $\mu$ g mL<sup>-1</sup>, respectively; and for FBZ they were 2.4575 and 8.1918  $\mu$ g mL<sup>-1</sup>, respectively. The analytical performance parameters followed the recommendations of the guide for the validation of analytical methods (ICH, 2005).

### Solubility studies

The phase solubility trials were performed according to the recommendations of Higuchi and Connors (1965). An excess of the drug (ABZ or FBZ) was added to dispersions containing increasing amounts of CDs (0 to 40 mmol L<sup>-1</sup>) in the absence or presence of 0.1% PVP-k-30 as a co-complexing agent. The dispersions obtained were stirred (DragonLab MS-MS10 10-position magnetic stirrer) and kept at equilibrium for 72 h at room temperature. After the equilibrium time had passed, aliquots of the supernatant were filtered through a 0.45  $\mu$ m cellulose acetate membrane filter, diluted with the mobile phase, transferred to a 1.5 mL amber vial, and the amount of drug dissolved was evaluated by liquid chromatography.

The intrinsic solubility of the drugs (ABZ and FBZ) and the solubility of the binary and ternary complexes in purified water were determined using similar conditions.

### Preparation of binary and ternary cyclodextrin complexes

The preparation of the binary (Drug/ $\beta$ CD or Drug/HP $\beta$ CD) and ternary (Drug/ $\beta$ CD-PVP 0.1% or Drug/HP $\beta$ CD-PVP 0.1%) complexes was performed by lyophilization: equimolar quantities of each drug, ABZ or FBZ, and CD,  $\beta$ CD, or HP $\beta$ CD, were weighed and dispersed, separately, in purified water. The obtained dispersions were submitted to homogenization (DragonLab MS-MS10 magnetic stirrer) for 2 h, transferred to borosilicate lyophilization vials with a round bottom ( $\phi$ , 70 mm/300 mL), frozen at -78.5 °C with dry ice, protected from light by aluminum foil, and lyophilized (Liotop Lyophilizer L101) for 120 h. After freeze-drying, the complexes were crushed in a glass mortar and pulverized in a stainless-steel particle sieve with openings of 0.71

mm. For the ternary complex, equimolar quantities of drugs and CDs were separately dispersed in a 0.1% PVP solution (w/v). The dispersion results were subjected to the previously described conditions for solubility trials.

### Thermogravimetric Analysis (TG) and Differential Scanning Calorimetry (DSC)

DSC curves were obtained by a Shimadzu Differential Scanning Calorimetry DSC-60 system using a sample mass of 2.0 to 4.0 mg in a previously calibrated aluminum capsule with metallic indium (99.99% purity;  $T_{\text{fusion}} = 156.4\text{ }^{\circ}\text{C}$ ;  $\Delta H_{\text{fusion}} = 28.7\text{ J g}^{-1}$ ), under a dynamic nitrogen atmosphere ( $100\text{ mL min}^{-1}$ ) with a heating rate of  $10\text{ }^{\circ}\text{C min}^{-1}$  from 40 to  $400\text{ }^{\circ}\text{C}$ .

Thermogravimetric curves (TG/DTG) were obtained by a Shimadzu differential thermogravimetry DTG-60 system using a sample mass of 4.0-6.0 mg in a previously calibrated aluminum capsule with calcium oxalate monohydrate (ASTM 1582-93) under a dynamic nitrogen atmosphere ( $100\text{ mL min}^{-1}$ ) with a heating rate of  $10\text{ }^{\circ}\text{C min}^{-1}$  from 40 to  $900\text{ }^{\circ}\text{C}$ . The obtained curves were adjusted with the TA-60WS software.

### Powder X-ray Diffraction (P-XRD)

The powder x-ray diffractograms were obtained with a Bruker diffractometer D8 Advance, using a Cu K $\alpha$  radiation source at 40 KV and 40 mA, in the range of  $5\text{-}60^{\circ}$  ( $2\theta$ ), with a scan rate of  $2^{\circ}\text{ min}^{-1}$ , using polycrystalline silicon (Si) as the default.

### Particle size and zeta potential measurements

The zeta potential and particle size/distribution measurements were performed on a Zetasizer Nano ZS/Malvern Instruments device. The scattering angle employed was  $173^{\circ}$ . The samples were diluted in milli-Q water at  $25\text{ }^{\circ}\text{C}$  in a 1:100 (w/v) ratio, submitted to an ultrasonic bath (Altasonic 9L) for 15 min, filtered through a  $45\text{ }\mu\text{m}$  syringe filter, and taken for analysis. The results are reported as the mean of three measurements.

### NMR spectroscopy

The  $^1\text{H}$  NMR spectra were obtained at 300 MHz (Bruker Avance, TXI SB 5 mm). The compounds were diluted to  $\sim 10\text{ mg mL}^{-1}$  in DMSO- $d_6$ . The chemical shifts were recorded in ppm ( $\delta$ ) using tetramethylsilane (TMS) as an internal reference.

### In vitro release studies

Dissolution trials were performed on an Ethik Technology 299-6/TS dissolution device, using a rotating disc apparatus (Pharmacopeia, 2016), 900 mL of ultrapure water as a dissolution medium, a stirring speed of 100 rpm, and a temperature of  $37 \pm 0.5\text{ }^{\circ}\text{C}$ .

For the preparation of the pills, amounts equivalent to 100 mg of the drug (ABZ or FBZ) and of the binary and ternary complexes were weighed and transferred to the matrices (rotating discs), and with the help of a 10-ton hydraulic press (Flowscience Instruments & Trade, Brazil), they were compacted under a pressure of 250 psi for 30 s. Subsequently, the matrices were coupled to the rotating disc apparatus and placed in the dissolution device. Aliquots of 10 mL were collected at time intervals of 5, 10, 15, 20, 30, 40, 50, 60, 90, 120, 180, and 240 min, and then placed back in the dissolution medium at the same temperature. The amount of dissolved drug was evaluated through liquid chromatography. The values obtained were corrected (Aronson, 1993). The accumulated amount of drug dissolved was plotted as a function of time.

## RESULTS AND DISCUSSION

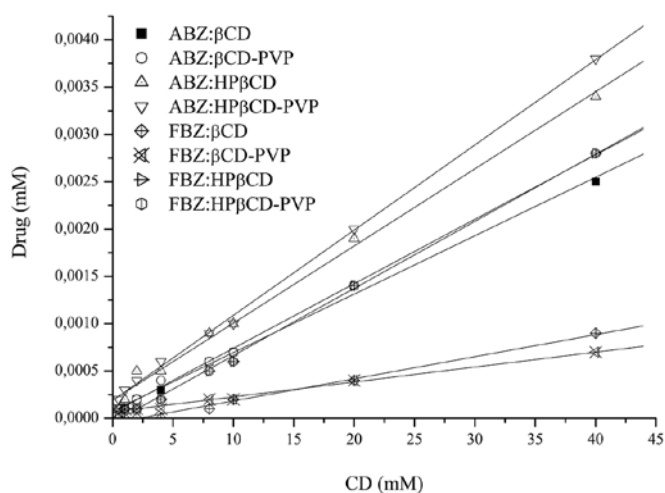
The main goal of adding CDs to pharmaceutical products is to increase the solubility of low-solubility drugs. The amount of CD needed to obtain the desired solubility is determined by the phase solubility diagram.

The phase solubility diagrams of the drugs ABZ and FBZ in the presence of  $\beta\text{CD}$  and  $\text{HP}\beta\text{CD}$  were of the type  $A_L$  (Figure 1). The linear regression of the data generated angular coefficient values ( $S$ ) of less than 1, assuming a stoichiometry of 1:1 (Table I) (Messner *et al.*, 2010). The intrinsic solubility values ( $S_0$ ) obtained were lower than those observed for the intercept solubility ( $S_{int}$ ), the latter being used to calculate the stability constant of the complexes ( $k_{1:1}$ ) (based on the equation of Higuchi and Connors, 1965):

$$k_{1:1} = \text{Slope}/S_{int} (1 - \text{Slope})$$

where,  $k_{1:1}$  is the stability constant for 1st order diagrams,  $S$  is the angular coefficient, and  $S_{int}$  the linear coefficient.

According to Pitha, Szente, and Greenberg (1983), only complexes with a stability constant  $k_c$  between 200 and  $5000\text{ M}^{-1}$  seem to have practical applications, given that the unstable complexes ( $k_c < 200\text{ M}^{-1}$ ) lead to premature release of the drug and the very stable complexes ( $k_c > 5000\text{ M}^{-1}$ ) result in delayed or incomplete release of the drug in the body. For many complexes,



**FIGURE 1** - Phase solubility diagram of the drugs (ABZ and FBZ) in the presence of the CDs.

values of  $k_c$  between 100 to 20,000  $M^{-1}$  can be found (Challa *et al.*, 2005).

The addition of water-soluble polymers to an aqueous solution of CD and drug is often linked to an increase in the stability constant of the complex ( $k_c$ ) and a reduction of the amount of CD required for incorporation of the drug (Loftsson, Másson, and Sigurjónsdóttir, 1999). The presence of polymers in drug/CD systems decreases the crystallinity of the drugs and has a synergistic effect on the solubilizing action of the CDs (Kurkov, Loftsson, 2013).

The ternary complexes with ABZ had  $k_{1:1}$  values larger than those of the binary complexes, but this effect was not observed for the FBZ ternary complexes. Since the  $k_{1:1}$  values are strongly affected by the solubility values ( $S_{int}$ ), distortions may be observed due to errors arising from the analytical method.

The solubility of the complexes increased compared to the pure drug (Table II). As expected, the modified cyclodextrin, hydroxypropyl- $\beta$ -cyclodextrin, proved to be more effective in dissolving the drugs than the natural cyclodextrin,  $\beta$ -cyclodextrin.

The formation of inclusion complexes in solution does not necessarily guarantee attainment and isolation of the complexes in the solid state, since these are the result of a set of chemical equilibria. The solid product may be a fine dispersion of the drug inside the CD or even a mixture consisting of inclusion complexes, non-complexed drugs, self-organized CD aggregates, and even CDs in the hydrated form.

Many techniques have been used to investigate the formation of inclusion complexes in the solid state, each having its own characteristics that not only allow detection of the formation of inclusion complexes, but also understanding of their chemical structure and the interactions that are established between the various components that make up the system.

Thermo-analytic techniques such as DSC and TG/DTG are widely used for the characterization of supramolecular complexes, such as inclusion complexes with CDs (Table III). The DSC curves of the physical mixtures for both drugs (ABZ or FBZ) had a thermal behavior that is representative of the sum of the substances. In the binary complexes, the event relating to the fusion of the drugs (ABZ and FBZ) was maintained, suggesting the presence of free drugs in the medium; the presence of the co-complexing agent PVP proved to be favorable to the amorphization of the drug/CD systems (Table III).

The success of the interaction between the guest molecule and the CD can be qualitatively assessed by comparing the properties of the complexes and the CDs,

**TABLE I** - Parameters obtained from the linear regression of the curves of the phase solubility diagram

	$S_0$	$S_{int}$	$S$	$R$	$k_{1:1}M^{-1}$
ABZ	1.5153E-6				
ABZ/ $\beta$ CD		7.7710E-5	6.1735E-5	0.9987	794
ABZ/ $\beta$ CD-PVP 0.1%		5.9415E-5	6.8125E-5	0.9993	1146
ABZ/HP $\beta$ CD		1.9162E-4	8.1486E-5	0.9967	425
ABZ/HP $\beta$ CD-PVP 0.1%		1.8777E-4	9.0034E-5	0.9994	480
FBZ	3.5202E-7				
FBZ/ $\beta$ CD		4.9647E-5	2.3359E-5	0.9955	471
FBZ/ $\beta$ CD-PVP 0.1%		6.9020E-5	1.5764E-5	0.9956	228
FBZ/HP $\beta$ CD		4.5531E-5	7.0927E-5	0.9990	1558
FBZ/HP $\beta$ CD-PVP 0.1%		3.9934E-5	5.3624E-5	0.9983	1343

\* $S_0$ , intrinsic solubility;  $S_{int}$ , linear coefficient;  $S$ , slope;  $R$ , correlation coefficient and  $k_{1:1}$ , stability constant

**TABLE II** - Solubility values obtained for binary and ternary complexes

Complexes	Solubility ( $\mu\text{g mL}^{-1}$ )				SD	IS
	1	2	3	Average		
ABZ	0.3865	0.4176	0.4522	0.4188	0.03	1
ABZ/ $\beta$ CD	87.7026	91.1609	101.5427	93.4687	7.20	223
ABZ/ $\beta$ CD-PVP	122.5868	128.7522	124.6638	125.3343	3.14	299
ABZ/H $\beta$ CD	433.8756	457.9660	437.3408	443.0608	13.02	1058
ABZ/HP $\beta$ CD-PVP	566.8074	596.6097	620.2490	591.2220	22.22	1412
FBZ	0.0811	0.1096	0.1255	0.1054	0.02	1
FBZ/ $\beta$ CD	41.1571	47.7081	47.8255	45.5636	3.82	432
FBZ/ $\beta$ CD-PVP	32.5536	40.4370	41.3538	38.1148	4.48	362
FBZ/H $\beta$ CD	153.5421	165.5211	159.0208	159.3616	6.00	1512
FBZ/HP $\beta$ CD-PVP	140.0309	152.3906	141.5664	144.6626	6.74	1373

\*SD, standard deviation; IS, increase of solubility (expressed in number of times)

or quantitatively assessed by evaluating the concentration of the substance in the inclusion complexes. An indirect method to evaluate the quantity of substance within the CDs is to compare the amount of water in the CD and in the complexes. In the formation of the complexes, the water molecules inside the cavity of the CD are replaced by the guest molecule. Through thermogravimetric analysis of the CDs and complexes between ambient temperature and a temperature below 100 °C, one can evaluate the release of water from the inside of the cavity

of the CD and its replacement by the guest molecule (Moriwaki *et al.*, 2008; Hădărugă, Hădărugă, Isengard, 2012). In this case, the binary and ternary complexes showed a reduction of the water bonded to the inside of the cavity of the CDs, suggesting the replacement of water molecules and the interaction of the guest molecule with the CDs (Table III).

The diffractogram revealed intense reflections at 7, 11.2°, 13.8°, 17.9°, and 22° (2 $\theta$ ) for ABZ; and at 6.6°, 11°, 13°, 25.9°, and 26.3° (2 $\theta$ ) for FBZ, which are

**TABLE III** - Thermal parameters obtained from the DSC and TG/DTG curves of the drugs (ABZ and FBZ) and binary and ternary complexes

	DSC		TG <sub>&lt;100°C</sub>	TG	
	T <sub>peak</sub> (°C)	$\Delta\text{H}$ (J g <sup>-1</sup> )	$\Delta\text{m}$ (%)	T <sub>initial</sub> (°C)	$\Delta\text{m}$ (%)
$\beta$ CD			13.726	29/166	14/84
HP $\beta$ CD			5.596	28/144	6/94
ABZ	199/213	-78/-3	0.044	200/225/322/379/464	4/8/15/32/19*
ABZ/ $\beta$ CD	186-200	-42	10.110	29/150/257	10/3/72
ABZ/ $\beta$ CD-PVP	187-200	-39	9.827	30/155/259/393	10/3/58/16
ABZ/HP $\beta$ CD	187-204	-28	4.508	32/145/255	5/2/85
ABZ/HP $\beta$ CD-PVP	187-203	-25	5.182	31/131/239/405	6/2/70/16
FBZ	214/226	-105/-11	0.061	204/230/274/390/538	4/7/12/44/10*
FBZ/ $\beta$ CD	206-219	-65	8.900	26/165/273	11/3/70
FBZ/ $\beta$ CD-PVP	204-221	-57	9.590	28/132/261	10/3/72
FBZ/HP $\beta$ CD	206-226	-34	5.124	29/137/260	5/2/83
FBZ/HP $\beta$ CD-PVP	204-229	-38	5.260	30/125/254	5/2/87

\*waste

characteristics of the crystallinity of the compounds.  $\beta$ CD revealed intense reflections at  $6.1^\circ$ ,  $8.9^\circ$ ,  $10.5^\circ$ ,  $12.4^\circ$ ,  $12.6^\circ$ , and  $15.3^\circ$  ( $2\theta$ ), which are characteristic of the crystallinity of the substance. HP $\beta$ CD and PVP had amorphous behavior. The physical mixtures of the drugs (ABZ or FBZ) with both CDs ( $\beta$ CD or HP $\beta$ CD) had reflections that were characteristic of each one of the isolated substances; the presence of PVP did not favor the amorphization of the mixtures. The binary and ternary complexes had a behavior similar to the physical mixtures, but with less intense reflections.

According to Pradines *et al.* (2014) the increase in the solubility of drugs in the presence of CDs cannot be attributed only to the formation of inclusion complexes, but also to the formation of self-organizing CD nanoclusters, which increase the solubility of the hydrophobic molecule through a micellar-type mechanism. The stability constant ( $k_c$ ) describes the result of all the effects of the CDs on the solubility of the drug. Not only inclusion complexes and non-inclusion complexes are formed in the aqueous CD solution, but also bonds with the hydrogen molecules of neighboring CDs, forming self-organizing CD aggregates with an approximate diameter of 100 nm (Messner *et al.*, 2010).

Table IV shows the mean hydrodynamic radius ( $R_H$ ), polydispersity index (PI), and zeta potential (Z) values of the drugs (ABZ and FBZ), the CDs ( $\beta$ CD

and HP $\beta$ CD), and the binary and ternary complexes. Although the modified CDs have functional groups on their surface that prevent self-clustering, self-organizing CD aggregates were observed in the binary and ternary systems for both CDs (Table IV). In the systems obtained with  $\beta$ CD, nanoclusters with a mean hydrodynamic radius of  $188.6 \pm 11.7$  nm were identified; in the systems obtained with HP $\beta$ CD, CD aggregates ( $180.9 \pm 4.5$  nm) and CD monomers ( $1.4 \pm 0.3$  nm) were observed. The results are in agreement with the literature (Coleman *et al.*, 1992; González-Gaitano *et al.*, 2002; Wu, Shen, He, 2006; Glisoni *et al.*, 2012).

Negative zeta potential (Z) values were obtained, indicating the stability of the complexes formed. The presence of the water-soluble polymer seems to favor steric repulsion, since the zeta potential (Z) values in ternary complexes were, in absolute value, larger than those of the corresponding binary complexes. All samples showed high polydispersity (PI > 0.08).

Nuclear magnetic resonance spectroscopy (NMR) is one of the most commonly used techniques for analysis of the stability, stoichiometry, and geometry of inclusion complexes, because it allows study of the interactions between the guest molecule and the CD directly.  $^1\text{H}$  NMR studies can be used to monitor changes in the chemical shifts of protons as the composition of the complexes is

**TABLE IV** - Mean values of hydrodynamic radius ( $R_H$ , nm), polydispersity index (IP) and zeta potential (Z, mV) of drugs, CDs and binary and ternary complexes (n=6)

	$D_H$ (nm)	I (%)	IP	Z (mV)
$\beta$ CD	$188.6 \pm 11.7$	100	0.627	
HP $\beta$ CD	$180.9 \pm 4.5$ $1.4 \pm 0.3$	88.8 11.2	0.457	
ABZ	$564.7 \pm 42.61$	100	1.00	$-32.5 \pm 5.14$
ABZ/ $\beta$ CD	$715.7 \pm 167.5$ $167.8 \pm 29.82$	82.4 17.6	0.568	$-27.1 \pm 4.80$
ABZ/ $\beta$ CD-PVP	$555.0 \pm 167.4$ $137.2 \pm 27.94$	85.9 11.9	0.555	$-33.8 \pm 5.29$
ABZ/HP $\beta$ CD	$871.4 \pm 195.8$	100	0.574	$-23.7 \pm 3.74$
ABZ/HP $\beta$ CD-PVP	$969.4 \pm 98.11$	100	0.177	$-24.1 \pm 4.55$
FBZ	$838.7 \pm 232.1$ $237.2 \pm 56.44$	79.8 20.2	0.586	$-34.6 \pm 4.41$
FBZ/ $\beta$ CD	$680.4 \pm 119.5$	100	0.496	$-27.8 \pm 4.70$
FBZ/ $\beta$ CD-PVP	$510.8 \pm 122.5$ $140.6 \pm 23.88$	84.5 15.5	0.549	$-29.7 \pm 4.87$
FBZ/HP $\beta$ CD	$491.4 \pm 191.9$	100	0.557	$-31.2 \pm 4.42$
FBZ/HP $\beta$ CD-PVP	$757.6 \pm 239.3$ $218.0 \pm 53.83$	77.3 20.5	0.568	$-33.5 \pm 4.91$

varied, and, as such, they provide information about the stoichiometry of the complexes and the dynamics of their formation. At the same time, when the drug is incorporated into the cavity of the CD, the hydrogen atoms located inside them (H-3 and H-5) undergo a significant chemical shift to higher value ranges, while the hydrogen atoms located on the outer surface of the CD (H-1, H-2, H-4, and H-6) are either not affected or undergo negligible shifts (Fernandes *et al.*, 2003; Ikeda *et al.*, 2004; Ferreira *et al.*, 2015).

Figure 2 shows the  $^1\text{H}$  NMR spectra of the benzimidazoles, ABZ (A) or FBZ (B), recorded at 300 MHz in  $\text{DMSO}-d_6$  with the signal assignments of the hydrogens. Figure 3 corresponds to the ABZ spectra in  $\beta\text{CD}$  binary (A) and ternary (B) systems. The spectrum recorded for ABZ in the presence of  $\beta\text{CD}$  (A) reveals the presence of signals corresponding to the drug and the cyclodextrin. The absence of protection, and the broadening of the signals in the range between 6 and 4 ppm, corresponding to OH groups and  $\text{CH}_2$  groups of the pyran rings present in the  $\beta\text{CD}$  structure, suggest there are intermolecular interactions between the drug and  $\beta\text{CD}$ , but not effective complexation (Moriwaki *et al.*, 2008; De Jaeghere *et al.*, 2013). The addition of PVP (B) caused changes in the profile of the hydroxyl groups at 5.70 ppm (broad singlet), suggesting a better interaction between ABZ and  $\beta\text{CD}$  in the presence of PVP.

The  $^1\text{H}$  NMR spectra of ABZ in the presence

of HP $\beta\text{CD}$  indicate the low solubility of the drug and the attempt to form inclusion complexes as shown in Figure 4 (A). The addition of the water-soluble polymer did not significantly alter the solubility of the system; no change in  $^1\text{H}$  NMR parameters was observed in Figure 4 (B).

Figure 5 shows the  $^1\text{H}$  NMR spectra of FBZ in the binary systems with  $\beta\text{CD}$  (A) and HP $\beta\text{CD}$  (B). The  $^1\text{H}$  NMR spectra of FBZ in the presence of both CDs revealed the presence of signals corresponding to the drug and the CDs, suggesting an interaction between the molecules. In the  $^1\text{H}$  NMR spectrum of FBZ in the presence of HP $\beta\text{CD}$  (B), enlargement and distortion of the signals in the region between 6 and to 4 ppm is visible, corresponding to the OH groups and  $\text{CH}_2$  groups of the pyran rings present in the HP $\beta\text{CD}$  structure, indicating effective complexing between FBZ and HP $\beta\text{CD}$ . The spectra recorded after the addition of PVP to the binary systems FBZ/CD ( $\beta\text{CD}$  or HP $\beta\text{CD}$ ) did not reveal any alterations in the NMR spectrum.

The great importance of CDs in the development of medication containing low-solubility drugs lies in their ability to act as the carrier of hydrophilic drugs, thereby increasing the solubility and/or dissolution rate, and, consequently, the bioavailability of these drugs (Challa *et al.*, 2005). The dissolution trials of the drugs and their complexes enabled the identification of differences in the dissolution rate of the binary and ternary complexes in

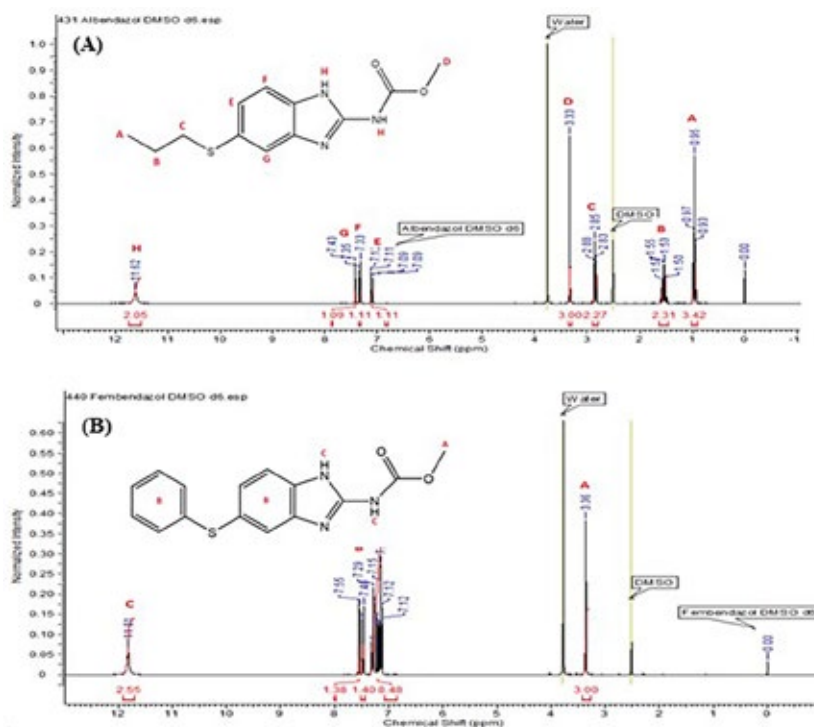
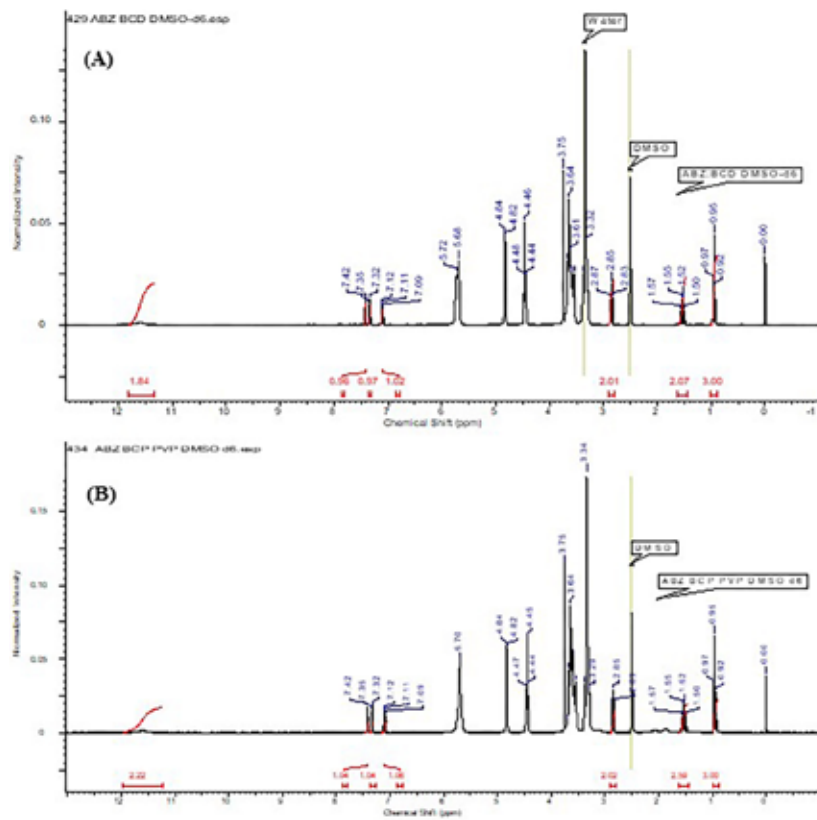
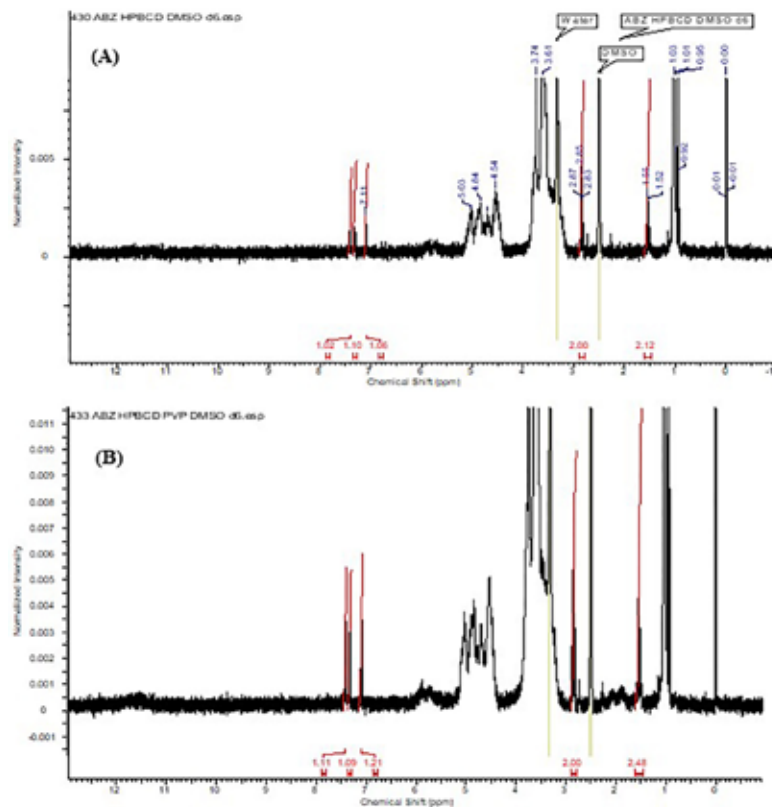


FIGURE 2 -  $^1\text{H}$ -NMR spectra of ABZ (A) and FBZ (B) recorded in 300 MHz  $\text{DMSO}-d_6$ .



**FIGURE 3** - <sup>1</sup>H-NMR spectra of ABZ in the presence of βCD (A) and the water-soluble polymer (PVP) (B) recorded in 300MHz DMSO-d<sub>6</sub>.



**FIGURE 4** - <sup>1</sup>H-NMR spectrum of ABZ in the presence of HPβCD (A) and the water-soluble polymer (PVP) (B) recorded in 300 MHz DMSO-d<sub>6</sub>.



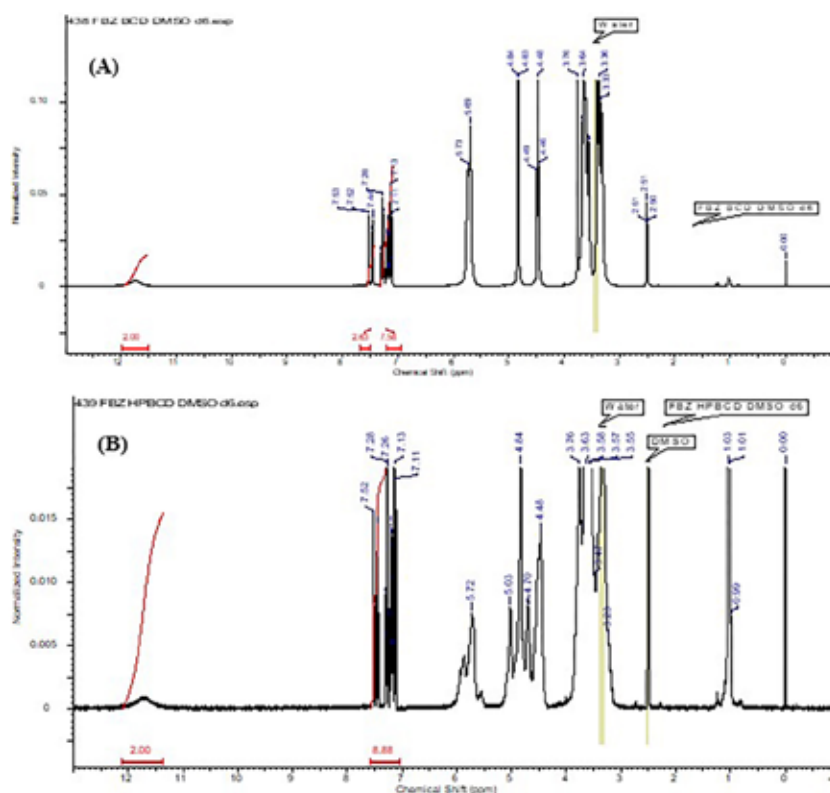


FIGURE 5 - <sup>1</sup>H-NMR spectra of FBZ in the presence of βCD (A) and HPβCD (B) recorded in 300 MHz DMSO-d<sub>6</sub>.

relation to the pure drugs. The dissolution rates of the pure drugs and of the binary and ternary complexes obtained with βCD were lower than that observed for HPβCD. In turn, the ternary complexes obtained with HPβCD had higher dissolution rates than the binary complexes (Figure 6). According to Kumar *et al.* (2011b) the increase in solubility and dissolution rate of FBZ in binary and ternary complexes obtained with β-cyclodextrin and

PVP-k-30 can be attributed to the increase in the drug's wetting properties.

## CONCLUSION

Among the CDs compared in this study to increase the solubility of the benzimidazole-2-carbamates albendazole and fenbendazole, hydroxypropyl-β-cyclodextrin showed the most potential. A synergistic effect between the water-soluble polymers PVP-k30 and HPβCD enabled a significant increase in the solubility and dissolution rate of the drugs.

The analysis of the changes in displacement and the profile of the bands matching the cyclodextrins in the <sup>1</sup>H NMR spectra revealed a molecular interaction and pointed to an effective complexation in the drug/CD systems.

The increase in the solubility of the drugs, either by the formation of inclusion complexes or by the formation of self-organizing CD aggregates, constitutes a promising strategy for the pharmaceutical industry to increase the solubility of these drugs. However, in addition to low solubility, ABZ and FBZ have other problems, such as poor stability and intestinal permeability. Therefore, the biological evaluation of these drugs and their complexes will be pursued in our future studies.

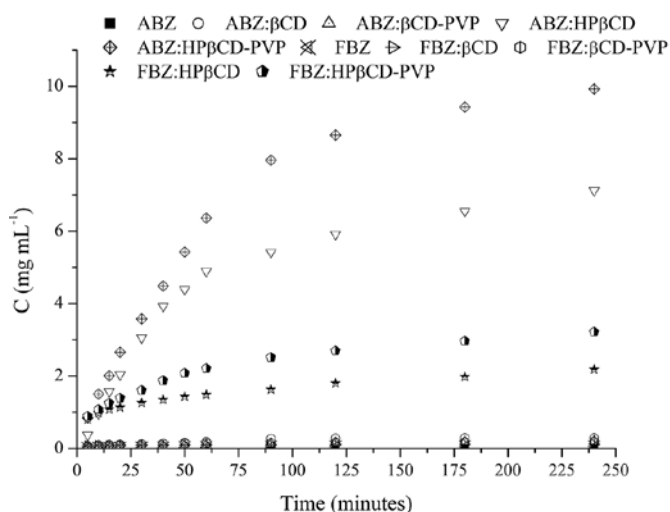


FIGURE 6 - Amount dissolved *versus* time obtained for the drugs (ABZ and FBZ) and their binary and ternary complexes.

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