

Effect of light sources and curing mode techniques on sorption, solubility and biaxial flexural strength of a composite resin

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ABSTRACT

Adequate polymerization plays an important role on the longevity of the composite resin restorations. Objectives: The aim of this study was to evaluate the effect of light-curing units, curing mode techniques and storage media on sorption, solubility and biaxial flexural strength (BFS) of a composite resin. Material and Methods: Two hundred and forty specimens were made of one composite resin (Esthet-X) in a stainless steel mold (2 mm x 8 mm Ø), and divided into 24 groups (n=10) established according to the 4 study factors: light-curing units: quartz tungsten halogen (QTH) lamp and light-emitting diodes (LED); energy densities: 16 J/cm² and 20 J/cm²; curing modes: conventional (CM) and pulse-delay (PD); and permeants: deionized water and 75% ethanol for 28 days. Sorption and solubility tests were performed according to ISO 4049:2000 specifications. All specimens were then tested for BFS according to ASTM F394-78 specification. Data were analyzed by three-way ANOVA followed by Tukey, Kruskal-Wallis and Mann-Whitney tests ($\alpha=0.05$). Results: In general, no significant differences were found regarding sorption, solubility or BFS means for the light-curing units and curing modes ($p>0.05$). Only LED unit using 16 J/cm² and PD using 10 s produced higher sorption and solubility values than QTH. Otherwise, using CM (16 J/cm²), LED produced lower values of BFS than QTH ($p<0.05$). 75% ethanol permeant produced higher values of sorption and solubility and lower values of BFS than water ($p<0.05$). Conclusion: Ethanol storage media produced more damage on composite resin than water. In general the LED and QTH curing units using 16 and 20 J/cm² by CM and PD curing modes produced no influence on the sorption, solubility or BFS of the tested resin.

Key words: Composite resins. Solubility. Solvents. Mechanical phenomena.

INTRODUCTION

Light-cured composite resins are being more used specially because they are aesthetically advantageous and enable to be polymerized by physical and chemical activators^{8,30}. However, the crucial point to be reached during restorative dentistry procedures with composite resins is to obtain satisfactory restorations with an adequate light activation technique²⁴.

Composites consist mainly of filler particles and a resin matrix based on different monomers⁶. Currently, the most widely marketed composite resins are based on the polymerization of bisphenol A glycol dimethacrylate (Bis-GMA) and urethane dimethacrylate (UDMA)²⁵. The curing process occurs in the organic matrix where there is a monomer-polymer conversion through an activation mechanism⁸. However, an insufficient degree of conversion directly affects the physical properties

and chemical stability of material⁸. Polymers used in the composition of composite resins may absorb water and chemicals from the oral environment (sorption) and may release some of their own components (solubility)^{1,12}. Restorations are subject to tension due to chewing when an excessive or continue non-functional load is applied, internal stresses and strain are generated which can cause the material failure²⁷.

In this context, the adequate polymerization plays an important role on the longevity of composite resin restorations. This procedure requires sufficient light energy intensity and an adequate wavelength in order to activate the photoinitiator within these materials, which will react with the reducer agent to form free radicals and initiate the polymerization process²⁴. Variables such as adequate light intensity, correct wavelength and energy density (power density x exposure time) are essential for achieving the proper depth of cure¹. Different types of light curing units have been proposed for the polymerization of light activated resinous materials including conventional quartz tungsten halogen (QTH) lamp, plasma arc, argon laser and light-emitting diodes (LED). All these devices emit blue light in the spectrum of the absorption of camphorquinone, which is the photoinitiator generally found in most resinous material⁹. The QTH lights can achieve satisfactory results, but they present certain limitations such as shorter durability of the lamp and heat production⁹. LED have been recognized as a promising technology for polymerization of resin-based materials because all the light emitted is within the spectrum of maximum absorption of camphorquinone at 468 nm^{1,13}.

Different curing regimes can exert an influence on the resistance of the union, hardness and crosslinking density⁷. When high power density is applied, more photoinitiator molecules are activated at the same time and consequently acceleration occurs immediately upon light exposure, thereby generating inner stresses that are confined into composite, dental structure and interfaces²¹.

In the pulse-delay (PD) polymerization technique there is a decrease in shrinkage and stress³⁰. In this technique, there is a short period (about 1-3 min) between initial exposition and final cure¹⁰. The interval between both pulses, an initial pulse with low energy density and a final pulse with high light irradiation, increases the composite resin pre-gel phase improving its flowability and relieving the stress generate by composite shrinkage. Thus, a better marginal seal is obtained^{17,30}. It has been hypothesized that low power densities, characterized by reduced power density in the initial seconds, may generate a small number of free radicals. Then, a more linear polymeric structure is obtained, with lower crosslinking density, which

negatively affects the material's chemical and physical properties²⁸. A previous study⁴ concluded that marginal integrity of restorations and the composite "plasticizing effect" are negatively affected when a PD technique is used.

The aim of this study was to investigate the effect of different light sources, PD curing protocols and storage media on sorption, solubility and biaxial flexural strength (BFS) of a composite resin. The null hypotheses tested were: (1) QTH and LED curing light units using different energy densities produce no influence on sorption, solubility and BFS; (2) QTH and LED curing light units using different curing techniques produce no influence on sorption, solubility and BFS; (3) storage procedure in water and ethanol produces no influence on sorption, solubility and BFS.

MATERIAL AND METHODS

Preparation of specimens

A nanohybrid composite resin (Esthet-X, Dentsply, Rio de Janeiro, RJ, Brazil) (Batch #071116), A2 shade, was used to perform the sorption and solubility test and biaxial piston-on-ring flexural strength test. The specimens were made in compliance with ISO 4049:2000 standard specifications¹⁴. Specimens were made using a stainless steel mold with 2 mm in thick and 8 mm in diameter. Two hundred and forty specimens were divided into 24 groups (n=10) established according to the 4 study factors:

1- light-curing units: QTH lamp (Variable Intensity Polymerizer, Bisco Inc., Schaumburg, IL, USA) and LED (Ultra Blue IS, DMC, São Carlos, SP, Brazil), 2- energy densities: 16 J/cm² and 20 J/cm², 3- curing modes: conventional mode (CM), pulse-delay using 3 s (PD3s) and pulse-delay using 10 s (PD10s); 4- storage media (for 28 days): deionized water and 75% ethanol (Figure 1).

The light irradiance for each curing mode was checked with the in-built radiometer prior to use to ensure consistency of light output. A single increment of composite material was placed into a stainless steel mold and confined between two opposing polyester strips (K-Dent, Quimidrol, Joinville, SC, Brazil). The material was irradiated according each protocol. The specimens were then removed from the mold and excesses were eliminated with a scalpel blade (Free-Bac, Wuxi Xinda Medical Device Co. Ltd., Wuxi City, Jiangsu, China).

Sorption and solubility tests

The sorption and solubility tests were performed in compliance with ISO 4049:2000¹⁴ standard specifications, except for the specimen dimensions, types of permeant and storage time. The specimen

was placed individually into an open glass bottle of 20 mm (Verallia; Saint-Gobain Vidros S.A., São Paulo, SP, Brazil) inserted into the desiccators (Vidrolabor; Vidrolabor Ind. Com., São Paulo, SP, Brazil) containing freshly dried white silica, batch #0506198, (Vetec, Vetec Química Fina Ltda, Rio de Janeiro, RJ, Brazil), and maintained for 22 h at 37±1°C in a vacuum oven (Model 440 D, LF Equipamentos, São Paulo, SP, Brazil). Thereafter, desiccators containing the specimens were removed from the oven and placed on a bench for 2 h, at a temperature of 23±1°C, completing a cycle of 24 h. Specimens were weighed daily on an analytical scale accurate to 0.001 mg (Marte AY220, Marte Balanças e Aparelhos de Precisão Ltda, Santa Rita do Sapucaí, MG, Brazil). The complete cycle was repeated until a constant mass (M1) was obtained, that is, until the mass loss of each specimen was not more than 0.1 mg per 24 h cycle. Afterwards, the specimens were returned to their respective labeled bottles, and 15 mL of either deionized water or 75% ethanol were added with manual pipettes. The bottles were capped, replaced in the oven and kept at 37±1°C for 28 days.

After this period, all bottles were removed from

the oven and kept at room temperature, 23±1°C for 2 h. The specimens were removed from the bottles, dried with absorbent paper for 15 s and left in a sterile bucket (Duflex, SSWhite, Rio de Janeiro, RJ, Brazil) for 1 min. Specimens were weighed to obtain M2. Then the specimens were returned to the desiccators until they reached a constant weight (M3) using the same procedure described to obtain M1.

The values for water sorption and solubility were calculated in µg/mm³ using the following equations: $Wsp=(M2-M3)/V$; $Wsl=(M1-M3)/V$, where *Wsp* was the sorption, *Wsl* was the solubility, *M1* was a initial mass of the sample in mg, *M2* was a mass of the specimen after immersion into storage media in mg, *M3* was a final mass of the specimen recorded after to evaporation of the water or ethanol in mg, and *V* was a volume of the specimen in mm³.

Biaxial flexural strength test

After sorption and solubility tests, all specimens were submitted to a biaxial piston-on-ring flexural strength test. The specimen was positioned on a circular metallic device with 3 mm in radius that contain 3 symmetrically spaced steel spheres.

Light sources	Power density (J/cm ²)	Curing modes	Curing profiles		
QTH	16	Conventional (CM)	27 s-600 mW/cm ²	___	___
		Pulse delay (PD3 s)	3 s-300 mW/cm ²	2 min	25 s-600 mW/cm ²
		Pulse delay (PD10 s)	10 s-300 mW/cm ²	2 min	22 s-600 mW/cm ²
	20	Conventional (CM)	34 s-600 mW/cm ²	___	___
		Pulse delay (PD3 s)	3 s-300 mW/cm ²	2 min	32 s-600 mW/cm ²
		Pulse delay (PD10 s)	10 s-300 mW/cm ²	2 min	28 s-600 mW/cm ²
LED	16	Conventional (CM)	27 s-600 mW/cm ²	___	___
		Pulse delay (PD3 s)	3 s-300 mW/cm ²	2 min	25 s-600 mW/cm ²
		Pulse delay (PD10 s)	10 s-300 mW/cm ²	2 min	22 s-600 mW/cm ²
	20	Conventional (CM)	34 s-600 mW/cm ²	___	___
		Pulse delay (PD3 s)	3 s-300 mW/cm ²	2 min	32 s-600 mW/cm ²
		Pulse delay (PD10 s)	10 s-300 mW/cm ²	2 min	28 s-600 mW/cm ²

Figure 1- Details of the various light curing modes evaluated.

Table 1- Means, standard deviations (±SD) for comparison between light source and permeant for the measurement of sorption and solubility at 16 J/cm²

Interation light source + permeant	Sorption			Solubility		
	CM	PD3 s	PD10 s	CM	PD3 s	PD10 s
QTH-water	9.97 ^a (0.95)	11.29 ^a (2.66)	10.76 ^a (1.88)	1.71 ^a (0.78)	3.08 ^a (2.40)	1.68 ^a (1.56)
LED-water	10.02 ^a (1.39)	12.08 ^a (1.67)	11.22 ^a (1.48)	1.68 ^a (1.48)	2.04 ^a (1.64)	3.01 ^a (1.45)
QTH-ethanol	19.02 ^b (2.21)	18.77 ^b (0.94)	18.39 ^b (2.60)	6.75 ^b (2.33)	6.33 ^b (2.07)	6.74 ^b (1.47)
LED-ethanol	19.98 ^b (1.46)	20.23 ^b (1.96)	21.54 ^c (2.56)	6.86 ^b (1.59)	8.47 ^b (1.86)	10.70 ^c (2.22)

Values in the column with same superscript lower caser letter denote no statistical differences (p>0.05)

Table 2- Means, standard deviations (\pm SD) for comparison between light source and permeant for the measurement of sorption and solubility at 20 J/cm²

Interaction light source + permeant	Sorption			Solubility		
	CM	PD3 s	PD10 s	CM	PD3 s	PD10 s
QTH-water	9.96 ^a (\pm 1.53)	10.49 ^a (\pm 2.34)	12.15 ^a (\pm 1.71)	0.36 ^a (\pm 0.61)	1.41 ^a (\pm 1.83)	1.78 ^a (\pm 1.28)
LED-water	10.24 ^a (\pm 1.77)	11.40 ^a (\pm 2.74)	10.84 ^a (\pm 0.94)	0.01 ^a (\pm 0.00)	3.02 ^a (\pm 2.71)	1.63 ^a (\pm 1.49)
QTH-ethanol	16.58 ^b (\pm 2.04)	18.93 ^b (\pm 2.20)	19.05 ^b (\pm 1.54)	4.95 ^b (\pm 2.16)	4.78 ^b (\pm 1.36)	6.45 ^b (\pm 1.88)
LED-ethanol	17.89 ^b (\pm 1.19)	20.52 ^b (\pm 2.53)	19.59 ^b (\pm 2.15)	4.95 ^b (\pm 1.53)	7.30 ^b (\pm 2.63)	6.55 ^b (\pm 2.54)

Values in the column with same superscript lower case letter denote no statistical differences ($p > 0.05$)

Table 3- Means, standard deviations (\pm SD) for biaxial flexural strength at 16 J/cm²

	QTH			LED		
	CM	PD3 s	PD10 s	CM	PD3 s	PD10 s
water	128.52 ^A (\pm 17.81)	107.09 ^B (\pm 21.17)	114.28 ^{A,B} (\pm 9.74)	106.62 ^B (\pm 24.97)	119.69 ^{A,B} (\pm 18.67)	122.15 ^{A,B} (\pm 15.14)
ethanol	63.09 ^C (\pm 7.67)	56.43 ^C (\pm 11.29)	51.76 ^C (\pm 5.88)	59.88 ^C (\pm 9.07)	50.69 ^C (\pm 6.85)	46.70 ^C (\pm 5.98)

Values with same superscript capital letter (A, B or C) denote no statistical differences ($p > 0.05$)

Table 4- Means, standard deviations (\pm SD) for biaxial flexural strength at 20 J/cm²

	QTH			LED		
	CM	PD3 s	PD10 s	CM	PD3 s	PD10 s
water	121.32 ^{Aa} (\pm 25.40)	111.98 ^{Aa} (\pm 23.79)	130.33 ^{Aa} (\pm 16.39)	118.09 ^{Aa} (\pm 28.90)	118.44 ^{Aa} (\pm 9.20)	111.26 ^{Aa} (\pm 15.14)
ethanol	68.36 ^{Ba} (\pm 11.44)	60.13 ^{Ba} (\pm 9.36)	58.21 ^{Ba} (\pm 10.13)	58.30 ^{Ba} (\pm 12.99)	55.85 ^{Ba} (\pm 6.94)	55.80 ^{Ba} (\pm 7.84)

Values in the same column with same superscript minute letter (A or B) denote no statistical differences ($p > 0.05$)

Values in the same line with same superscript capital letter (a or b) denote no statistical differences ($p > 0.05$)

Each specimen was placed concentrically on the supporting spheres to ensure that the load was applied at the center of the specimen. A steel piston of 2 mm in diameter was attached to a universal testing machine (EMIC, DL2000, São José dos Pinhais, PR, Brazil) and the test was performed at a crosshead speed of 0.5 mm/min using a 50 Kg load cell until specimen failure. The BFS was calculated according to the ASTM F394-78 specifications as follows²:

$S = -0.2387 P(X-Y)/d^2$; where S is the flexural strength in MPa, P is the load at failure in N, and d is the specimen thickness in mm of the failure area. X and Y were determined as follows: $X = (1+v) \ln(B/C)^2 + [(1-v)/2] (B/C)^2$ and $Y = (1+v) [1 + \ln(A/C)^2] + (1-v) (A/C)^2$, where v is Poisson's ratio of the composite resin (0.24)²⁰, A is the radius of the support circle, B is the radius of the tip of the piston, and C is the radius of the specimen.

Data obtained from the sorption, solubility and BFS tests were subjected to three-way analysis of

variance and Tukey's tests for parametric values, and Kruskal-Wallis and Mann-Whitney tests for non-parametric values ($\alpha = 0.05$). Pearson's correlation was used to verify the correlation among the sorption, solubility and BFS. All tests were performed using the SPSS 17.0 for Windows statistics software (SPSS Inc., Chicago, IL, USA).

RESULTS

The non-parametric data were analyzed in terms of the average position. However, to facilitate understanding, the data presented in the tables are the means of the groups with their respective standard deviations (\pm s.d.). The results for sorption and solubility at energy densities of 16 J/cm² and 20 J/cm² are shown in Tables 1 and 2, respectively. In general, no significant difference was found between the light sources and between the curing modes tested in this study ($p > 0.05$). However, in the interaction between permeant and light source

Table 5- Pearson's correlation coefficient

	Sorption	Solubility	BFS
Sorption	---	+ 0.852*	- 0.840*
Solubility	+ 0.852*	---	- 0.781*
BFS	- 0.840*	- 0.781*	---

* significant correlation ($p < 0.05$).

with a density of 16 J/cm², it was seen that the specimens light-cured with the LED unit device in PD10s and stored in ethanol led to higher values of sorption and solubility than those light-cured with QTH ($p < 0.05$). Ethanol showed higher sorption and solubility than water irrespective of curing unit or curing method ($p < 0.05$).

In the BFS there were interactions between light sources, curing modes and permeants at an energy density of 16 J/cm² ($p < 0.05$) (Table 3), which did not occur at a density of 20 J/cm², where there were significant differences only between the permeants (water and ethanol) (Table 4).

The correlation among sorption, solubility and BFS was performed by Pearson's correlation coefficient test. It was found that sorption and solubility data had a positive correlation. Otherwise, BFS were negatively correlated with sorption and solubility (Table 5).

DISCUSSION

Successful composite resin restoration depends on the association of low rate shrinkage, good flowability, appropriate cure and satisfactory mechanical properties¹³. LED sources seem to be a promising technology for polymerization of dental resin materials, but its application in composite curing remains controversial¹⁹. In this study, QTH lamp was adopted as a control in order to compare its performance with LED. LED units produce a narrow band of wavelengths (450-490 nm), conveniently situated within the absorption spectrum of camphorquinone, which is the photoinitiator present in the most of light-activated dental materials¹³. In general, the LED unit showed similar behavior to the QTH unit, suggesting that similar polymerization quality was reached by the equipment, except when LED using 16 J/cm² in PD10 s intensity produced higher values of sorption and solubility than QTH ($p < 0.05$). LED curing unit using CM at 16 J/cm² produced lower values of BFS than QTH used with the same protocol ($p < 0.05$). This fact could be associated with the greater heat generated by the QTH unit, which may speed up the polymer chain induction process in composite, increasing the mobility of molecules during the reaction and allowing more monomers to react before the curing process ends¹³. The polymerization does not stop

right after the photoactivation period, and heat contributes for this post-activation polymerization¹⁹. This could be further explained by the absorbance of camphorquinone. Although camphorquinone presents its maximum absorbance at 468 nm, it has an absorption band (380-510 nm) that is coincident with the light band emitted by QTH lamps resulting in more camphorquinone molecules being activated²². Consequently, the crosslinked polymer was probably lower in these LED subgroups than in the corresponding QTH subgroups and more susceptible to softening in solvents.

The size of the specimens used in this study differed from the ISO 4049:2000¹⁴, as the standard dictates that they must be 15 mm in diameter. This contrasts with the diameter of the light guide tip of the curing units, which are 8 to 10 mm in diameter, and which would limit the uniform irradiation throughout the circumference of the specimen.

Traditionally, the manufacturers have recommended high light intensity or power density to provide a higher degree of monomer conversion into polymer, thereby improving the mechanical properties of composite resin^{17,18}. High power density used for short light exposures led to a lower degree of cure and lower flexural strength and modulus than when the composite was cured with intermediary power densities for longer exposures²³. In some studies^{18,28} that used 24 J/cm² of energy density, the curing modes tested presented similar behavior. It has been speculated that this occurred as a result of the formation of a densely crosslinked polymer network because an adequate energy density had been used¹⁸. The same finding was seen in this study in the groups that used 20 J/cm², where there were no significant differences between the subgroups studied. The subgroups that used 16 J/cm² presented different values of sorption, solubility and BFS storage in ethanol for the QTH and LED units ($p < 0.05$). Consequently, the first hypothesis of this study was partially rejected.

The use of low initial intensity followed by high intensity light with an interval between them seems to create a uniform polymerization of the composite resin, providing its best adaptation to cavity walls and possibly the least polymerization contraction stress^{10,17,30}. In the delay period, little amount of free radicals and double bond conversion are produced and the composite resin has more time to molecular rearrangement and stress relief¹⁰. In the PD cure, there was reduced gap formation without any mechanical properties being compromised, which can be considered as an indicative that composite has the same quantity of remaining double bonds³. The use of slow-cure methods (pulse curing modes) in combination with the interval between two irradiations seem not to interfere with solvent sorption and could be useful

for adhesive composite restorations¹⁸. In this study, there were no significant differences among the curing modes in terms of sorption, solubility and BFS ($p > 0.05$) except for QTH comparing the CM and PD3s at 16 J/cm², where PD3s produced lower values of BFS than the CM ($p < 0.05$). Therefore, the second hypothesis proposed in the study was accepted. However, this result was different from those of other studies^{3,29}, where the PD mode resulted in a linear polymer structure that was less crosslinked, thereby increasing the susceptibility of polymers to softening in ethanol. Some studies^{3,16} have shown a significant reduction in mechanical properties, such as flexural modulus and hardness, when composites were photo-cured by PD mode and submitted to solvent action. However, based on the methodology used in this study and the obtained results, the curing modes presented similar behavior. According to Lopes, et al.¹⁸ (2009), this could have occurred because another study³ used unfilled resins for their tests instead of filled composite resin. Another explanation for the difference between the results of this study and that where the photo-cured composites using the PD mode were more susceptible to softening in solvent could be attributed to the type of tests applied to assess the amount of crosslinking, such as hardness deterioration, degree of conversion and remaining double bonds, which only make an indirect assessment of the quality of the polymer³.

The effect of the different solvents was clearly significant in all groups; ethanol resulted in more solubility, sorption and lower BFS than water media. Therefore, the third hypothesis was rejected. In the ISO 4049:2000 specifications¹⁴, the storage time is 7 days, but in the present study the storage time of 28 days in the same way as established for other studies^{12,25} in which most of the composites studied reached saturation within 7-60 days. The amount of solvent uptake by the polymer is determined by differences in solubility parameters between the polymer and the permeant¹⁹. Water or solvent uptake into the resin phase of CR causes two opposing processes: the solvent will extract unreacted components, mainly monomer, thereby resulting in shrinkage, loss of weight and reduction in mechanical properties; conversely, solvent uptake leads to a swelling of the composite resin and an increase in weight²⁵. These phenomena of sorption and solubility may serve as precursors to a variety of chemical and physical processes which create biological concerns and produce deleterious effects on the structure and function of the polymer material¹².

It is desirable that composites have all monomers converted into polymers during the polymerization reaction and remain stable for long time¹¹. It may be assumed that, whereas the increase in

the degree of conversion reduced the solubility because the amount of unreacted monomers available for leaching out was lower due to the high percentage of reacted aliphatic C=C bonds from the dimethacrylate monomers⁶. Most dental polymers present ionic functional groups with water affinity, which in turn optimizes their hydrophilicity¹¹. The susceptibility for more linear or less crosslinked polymer to softening in solvents may be explained by the solvent-polymer interaction and, consequently, by Hoy's solubility parameter for polar forces¹⁶. Solvents that can form strong secondary bonds with the polymer chains can penetrate and replace the interchange secondary bonds, and thereby pull apart and dissolve linear and branched polymers²⁵. Distilled water, a solvent indicated in ISO specifications¹⁴ for resin-based filling materials, simulates the wet intraoral environment provided by saliva and water. However, intraoral conditions are clearly more complex than those achieved with distilled water in the laboratory²⁵. According to the US Food and Drug Administration Guidelines, a 75% ethanol-water solution is a clinically relevant food/oral simulating liquid^{25,26}. The 75% ethanol solvent is the most frequently used to simulate accelerated ageing of restorations as it has a solubility parameter matching that of BisGMA^{12,25}. When ethanol penetrates the polymer network, it causes the structure swell, thereby allowing for the release of uncured monomers and causing the dissolution of linear polymer chains¹². This expansion is facilitated when crosslink density is low, based on the fact that more space and pathways are available for solvent molecules to diffuse within the structure⁸. For this reason, water and 75% ethanol were used in this experiment.

The properties of composite resins are affected by the degree of conversion, but an analysis of the amount of crosslinking in the resins may provide a closer correlation to mechanical properties¹². Water sorption may deteriorate polymer mechanical properties and because of this the resin specimens were submitted to a mechanical test after the sorption and solubility tests. Tensile strength is generally considered to be the most meaningful property of these brittle materials in assessing the potential for failure of dental restorations⁵. Flexural strength can be measured using a three-point bending flexural test, a four-point bending flexural test or a biaxial flexure test; the drawback of the three-point and four-point flexural tests is the inherent sensitivity to flaws and defect near specimen edges¹⁵. As Ban and Anusavice⁵ (1990) related, the biaxial flexure test is recognized as a reliable technique since the maximum tensile stress occurs within the central loading area and edge failures are eliminated.

In the evaluation of the correlation among the

variables (sorption, solubility and BFS), it was possible to see how much higher the sorption and solubility were and how much lower the BFS ($p < 0.05$) was. The degree of conversion and crosslink density can influence the chemical degradation that is usually caused by oxidation and hydrolysis processes¹², and consequently reduce the physical and mechanical properties. Thus, polymers with lower crosslink densities are expected to undergo more softening in solvents²⁹, which results in greater sorption and solubility values but lower BFS values.

CONCLUSION

Within the limitations of this *in vitro* study methodology, it was concluded that, in general, the light sources (QTH and LED) and curing modes (CM and PD) did not influence the sorption, solubility or BFS of the tested resin. However, the different solvents (water and ethanol) did influence its sorption, solubility and BFS behavior.

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