

Facile Method to Tune the Particle Size and Thermal Stability of Magnetite Nanoparticles

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Nucleation and growth mechanism of iron oxide nanoparticles on zeolite template and their stability dependence are reported. Hyperfine field resulting from the variation of particle size indicates the effect of zeolite on particles growth; particle size decreases at lower concentration of zeolite. At higher concentration, a fraction of nano Fe₃O₄ experiences hyperfine field (45 and 49 T) similar to bulk particles. Effect of incubation and digestion time on the particles growth and the binding effect with zeolite are discussed. Annealing treatments show that the binding of nanoparticles with zeolite stabilizes the nanoparticles with regard to agglomeration and structural transformation. Thermogravimetry-differential thermal analysis (TG-DTA) shows that increase in dehydration temperature from 335.1 to 351.7 K results in zeolite content increasing from 0 to 1000 mg. Weight loss of the particles prepared in incubation time of 0.5 min is 9.46% and reaches 13.9% in 240 min. The weight loss remains practically constant at ca. 9% irrespective of the digestion method.

Keywords: magnetite, zeolite, confined growth, thermal stability

Introduction

Magnetic iron oxide (Fe₃O₄ and γ -Fe₂O₃) nanoparticles have significant importance in magnetic and catalytic industries. It is reported to have been used in miniaturization of devices, drug delivery, hyperthermia, heavy ion separations and batteries, etc.¹⁻⁴ Various methods such as co-precipitation, sol gel and solvothermal, etc., have been adopted to synthesize magnetic nanoparticles.⁵⁻⁷ Large particle size distribution and low magnetism of nanoparticles are of concern for applications. In parallel to the functionality of the iron oxide nanoparticles, stability is an important issue. Agglomeration hinders the functionality of the nanoparticles.⁸ Thermal treatment of the nanoparticles results in the magnetic and structural transformation. This restricts their application at elevated temperature ambience. To improve the desired properties, nanoparticles were synthesized on a template, which possesses uniform and monodisperse fine pores.⁹ Various templates exist for synthesis of nanoparticles. In the present

study, zeolite 13x has been used as it has uniform pores with specific charge affinity sites, which act as nucleating centers of the nanoparticles.¹⁰ Tetrahedral groups of SiO₄ and AlO₄ are the basic units with Si and Al at the center of the tetrahedron. Each AlO₄ tetrahedral group has an uncompensated charge and therefore a positive ion such as Na⁺, Ca²⁺, etc., occupies site near AlO₄ in the course of formation. The stoichiometric formula for zeolite 13x¹⁰ is Na₈₆[(AlO₂)₈₆(SiO₂)₁₀₆].264H₂O. It consists of two major cages known as β -cages and α -cages. These cages act as the host for the external agents like H₂O, transition ions and nanoparticles.

Zeolite, being a nontoxic material, is often used as carrier of drugs, e.g., doxorubicin.¹¹ It has been a promising candidate for bioengineering and catalysis technology.¹²⁻¹⁴ Zeolites doped with magnetic elements/compounds have also been an important species for various technological applications.¹⁵⁻¹⁷ Therefore, the iron oxide/zeolite will lead to a multifunctional material which can be used in both biological and non-biological applications. Transition metals and oxides were grown on zeolite using various techniques, such as wet impregnation, chemical

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vaporization and physical methods.¹⁸⁻²¹ The metals or oxides grown in zeolite using wet chemical methods occupy the pores.²² Zeolite can uptake up to 2 wt.% of the transition metal inside the network without significant modification of the structure.²² Numerous studies have been reported on zeolite incorporated with Fe atoms, but studies on the growth of iron oxide on zeolite and the functionality of the composites could not be found. The thermal stability of the iron oxide on the phase transformation has been of scientific interest.²³⁻²⁵ Stability has been improved by doping of transition metals and capping with organic and ceramic materials.²⁶ Capping prevents the surface interaction with targets leading to reduction in the catalytic properties. Bare and stable nanoparticles with tuned particle size are necessary for wide range of applications.

In the present study, composites of iron oxides and zeolites have been synthesized using wet chemical method. The structural, thermal and magnetic properties have been studied by Mössbauer spectroscopy, X-ray diffraction, thermo gravimetry-calorimetry, and transmission electron microscopy.

Experimental

Material synthesis

The synthesis of nanocomposites of magnetic iron oxide and zeolite 13x were carried out through a modified wet chemical method. An aliquot of 20 mL of $\text{FeSO}_4 \cdot 7\text{H}_2\text{O}$ (0.2 mol L^{-1}) and 20 mL of $\text{FeCl}_3 \cdot 6\text{H}_2\text{O}$ (0.4 mol L^{-1}) were mixed under vigorous stirring at 90°C followed by addition of NaOH (2 mol L^{-1}) solution for precipitation of Fe_3O_4 nanoparticles. The pH of the solution was maintained higher than 10 with rapid addition of NaOH so as to decrease the possibility of hydroxide formation.²⁷ The precipitated magnetic iron oxide nanoparticles were separated using a permanent magnet and washed several times with water and acetone for removal of inorganic and organic impurities, respectively. The composites were prepared in three different ways by varying (i) the concentration of zeolite 13x, (ii) the incubation time of zeolite in the precursor solution before precipitation of iron oxide and (iii) controlling the effective digestion time of nucleated iron oxide by tuning the addition time of zeolite after precipitation. Each of these methods is described in detail below.

(i) Variation of zeolite 13x concentration

In order to observe the effectiveness of zeolite concentration on the size and magnetic properties of iron oxide nanoparticles, different concentrations of dehydrated

zeolite 13x (annealed at 523 K for 2 h) were mixed in the precursor solution at 320 rpm for 15 min. These samples are denoted as MxZ (where x represents the amount of zeolite 13x added in mg).

(ii) Variation of incubation time

Incubation time is one of the key parameters to control the particle size and morphology. The incubation time of zeolite in precursor solution was varied as 0.5, 60 and 240 min before addition of NaOH to precipitate Fe_3O_4 nanoparticles. The concentration of zeolite 13x was maintained at 150 mg of zeolite and the digestion time after precipitation was maintained at 30 min. The samples prepared in this process were denoted as Ix (x: time in min).

(iii) Variation of digestion time

Controlling the growth of nanoparticles by addition of capping agent is a standard process.²⁸ Zeolite, being a cationic affinity material, has a tendency to restrict the growth of freshly nucleated Fe_3O_4 nanoparticles. In this process, the zeolite 13x was added to the precipitated solution after 0.5, 5 and 30 min of NaOH addition. The concentration of zeolite was fixed at 150 mg. The samples prepared in this process are denoted as Dx (x: time in min).

Characterization techniques

X-Ray diffraction (XRD) was performed using RIGAKU Ultima IV diffractometer with Cu-K α as radiation source and Ni-filtered with cross beam optics (CBO) monochromator; operating voltage 45 kV at 15 mA. The measurements were performed in steps of 0.05° and the scan rate of 2° min^{-1} . Transmission electron microscopy (TEM) images were obtained using a microscope (JEOL, model 1011) to determine the morphology, the mean particle size and the size distribution. The Mössbauer spectra were recorded with a constant acceleration transducer coupled to ^{57}Co in Rh matrix source with an initial activity of 25 mCi in the standard transmission geometry. The velocity calibration *per* channel was done with a $1.9 \text{ mg per } ^{57}\text{Fe cm}^2$ iron foil. Thermogravimetry (TG)/derivative thermogravimetry (DTG) and differential thermal analysis (DTA) experiments were performed in a thermo-balance Shimadzu DTG-60H. TG/DTG curves of samples were carried out under oxidant atmosphere (30 mL min^{-1}) using a Pt pan. For all samples, the mass used was approximately 7 mg, and a heating rate of 10 K min^{-1} was employed between 298-1073 K. The dehydration and phase transformation temperatures were determined by DTG and the heats of exothermic and endothermic processes were determined by DTA.

Results and Discussion

XRD patterns (Figure 1) confirm the nanoparticles to be cubic iron oxide ($\text{Fe}_3\text{O}_4/\gamma\text{-Fe}_2\text{O}_3$). The average particles size of M0Z (pure Fe_3O_4) was calculated to be 9 nm using Scherer formula. At lower concentration of zeolite 13x, a decrease in the Fe_3O_4 particle size (M25Z: 6 nm, M50Z: 6.5 nm and M75Z: 6 nm) is observed which increases to 10.6 nm in the presence of 100 mg of zeolite. On further increasing the zeolite concentration to 180 (M180Z) and 200 mg (M200Z), the particle size decreases to 8.5 and 3 nm, respectively. Due to sharp decrease in particle size, the peaks corresponding to the diffraction pattern from crystal planes get broader with increase in zeolite concentration. The lattice parameter of the pure Fe_3O_4 is 0.836 nm in agreement with the reported value of Fe_3O_4 (0.838 nm).²⁵ Zeolite phase was observed along with iron oxide only for sample M1000Z (1000 mg of zeolite). Low intensity diffraction peak from the (642) crystal plane of zeolite was observed in sample M500Z which sharply increases for sample M1000Z and dominates over the (533) peak which happens to be most intense peak of pure zeolite. It has been attributed to the incorporation of iron oxide nanoparticles which affects the crystal structure of zeolite. The XRD peaks of Fe_3O_4 after addition of zeolite is due to superposition of small size nanoparticle nucleated in the pores and the bigger particles residing on the surface, resulting into broad basal and sharp peaks.

Figure 1B shows the XRD pattern of the samples nucleated after 0.5 min of incubation (I0.5) and samples synthesized with 30 min of digestion time (D30). The diffraction pattern confirms presence of pure spinel iron oxides. Sample I0.5 shows the presence of weak and broad features of zeolite whereas sample D30 shows presence of sharp crystalline peaks of zeolite. This can be attributed to the amorphization of zeolite due to nucleation of the nanoparticles from the uncompensated cationic affinity sites of the zeolite.²⁹ However, the crystal structure of zeolite remains unaffected when the nanoparticles were nucleated and adsorbed on the zeolite surface. The amorphization of the zeolite is attributed to the strain developed during nucleation of nanoparticles.

Figure 2a shows the TEM image of the sample M0Z confirming the near to spherical shape of the nanoparticles with high monodispersity. Particle size distributions obtained from the TEM image analysis of M0Z shows particles are 14% of 4-8 nm, 54% of 8-12 nm and 32% of 12-20 nm. Figures 2c and 2d represent the high resolution (HR) TEM of M250Z and M1000Z, respectively.

Regular atomic patterns with interplanar spacing 0.25 nm representing the (311) plane of Fe_3O_4 and diameter

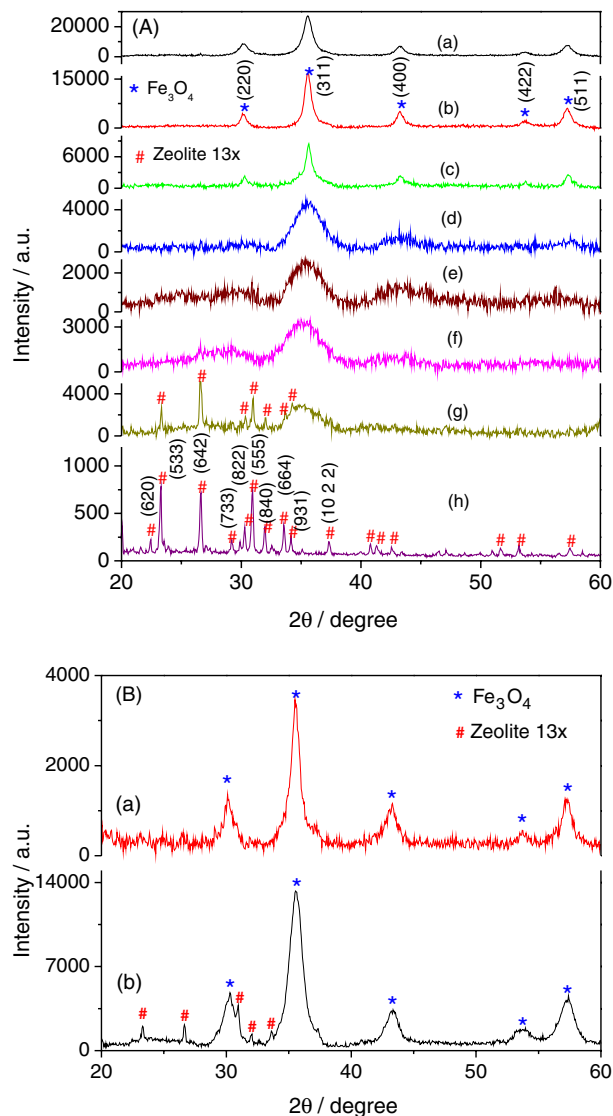


Figure 1. (A) X-ray diffractograms of (a) pure Fe_3O_4 (M0Z); (b) 100 mg (M100Z); (c) 180 mg (M180Z); (d) 200 mg (M200Z); (e) 250 mg (M250Z); (f) 500 mg (M500Z); (g) 1000 mg (M1000Z) of zeolite 13x with Fe_3O_4 and (h) zeolite 13x. (B) XRD pattern of the nanoparticles grown in the presence of 150 mg of dehydrated zeolite 13x with (a) incubation time of zeolite with solution of 0.5 min (I0.5); and (b) the 30 min digestion time of the iron oxide precipitate before adding zeolite (D30).

of 3 to 7 nm were observed. These regular interplanar spacing were absent in M1000Z. This is due to the reduction of Fe_3O_4 particles on the surface of zeolite. The XRD pattern observed in M1000Z is understood from the particles inside the pores.

Figure 3 shows the Mössbauer spectra of sample M0Z least square fitted to four sextets and a doublet, and Table 1 depicts the least square fitted Mössbauer parameters. The doublet represents the superparamagnetic particles (4 to 8 nm) which is 7% of the total area. The sextets (hyperfine field of 14.3 and 34.8 T) with area 61% can be assigned to the particles in the range of 8 to 12 nm. The remaining two

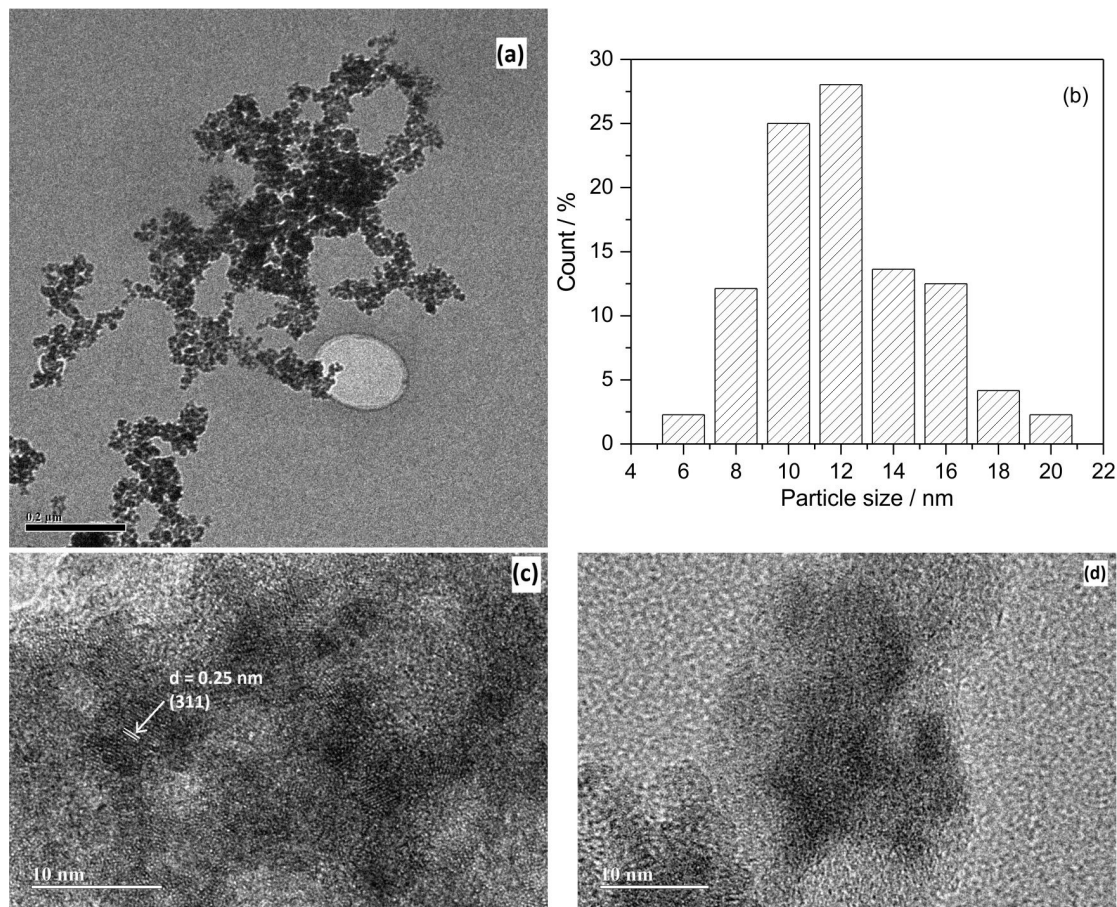


Figure 2. TEM image of (a) M0Z, (b) particles size distribution, (c) HRTEM of M250Z and (d) M1000Z.

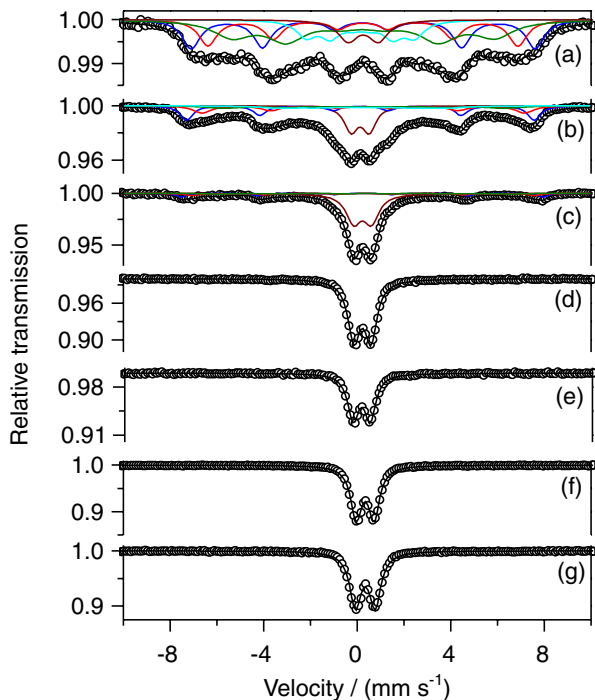


Figure 3. Room temperature Mössbauer spectra of (a) pure Fe₃O₄ (M0Z); composites of Fe₃O₄ and (b) 100 mg (M100Z); (c) 150 mg (M180Z); (d) 200 mg (M200Z); (e) 250 mg (M250Z); (f) 500 mg (M500Z) and (g) 1000 mg (M1000Z) of zeolite 13x.

sextet subspectra with internal magnetic fields of 45.6 and 41.1 T represent the Fe atoms of tetrahedral and octahedral sites of spinel Fe₃O₄ in the size range of 12 to 20 nm. The percentages of area fraction of different particle size from Mössbauer analyses (Table 1) are in good agreement with TEM analysis.

It is reported that zeolite possesses selective adsorption due to its unique porous structure and presence of uncompensated charge (AlO₄ units).³⁰ Dissolution of zeolite in the precursor solution results in adsorption of Fe²⁺/Fe³⁺ into the α and β cages with the excess Fe²⁺/Fe³⁺ ions remaining freely in the solution. The growth of the nanoparticles nucleated from the Fe²⁺/Fe³⁺ ions adsorbed inside the pores were restricted to the size of cage (maximum pore size in zeolite 13x is 13 Å) which results in the superparamagnetic doublet in Mössbauer spectra. The Fe₃O₄ nanoparticles nucleated from the freely remaining Fe²⁺/Fe³⁺ ions in solution undergo agglomeration which is reflected from sextet components.

The superparamagnetic doublet increases with the increase of zeolite which is attributed to the higher adsorption of Fe²⁺/Fe³⁺ ions into the pores (α and β cages) of zeolite. The Mössbauer spectra of samples

Table 1. Least square fitted Mössbauer parameters of samples M0Z, M100Z, M180Z, M200Z, M250Z, M500Z and M1000Z

Sample	Particle size / nm	IS (Fe) / (mm s ⁻¹)	QS / (mm s ⁻¹)	B _{hf} / T	Area / %	Phase ^a
M0Z	9.0	0.36(3)	1.30(0)	–	07(1)	Fe _{SPM}
		0.27(3)	0.04(4)	14.3(3)	12(2)	Fe _{Surf}
		0.34(2)	0.05(3)	34.8(5)	49(4)	Fe _{Surf}
		0.36(1)	0.01(2)	41.1(3)	15(4)	Fe _{Oct}
		0.34(1)	0.04(2)	45.6(2)	17(2)	Fe _{Tet}
M100Z	10.6	0.23(0)	0.77(1)	–	13(0)	Fe _{SPM}
		0.28(1)	1.54(0)	–	23(0)	Fe _{SPM}
		0.33(0)	0.09(0)	22.0(3)	19(0)	Fe _{Surf}
		0.41(3)	0.00(4)	37.2(5)	17(2)	Fe _{Surf}
		0.38(1)	0.04(2)	42.8(1)	14(2)	Fe _{Oct}
M180Z	8.7	0.24(7)	0.00(1)	46.1(0)	14(1)	Fe _{Tet}
		0.33(0)	0.76(0)	–	75(0)	Fe _{SPM}
		0.45(5)	0.03(0)	39.0(0)	08(1)	Fe _{Surf}
		0.40(2)	0.03(4)	45.3(3)	10(2)	Fe _{Oct}
M200Z	3.1	0.34(1)	0.03(3)	47.9(1)	07(1)	Fe _{Tet}
		0.33(0)	0.72(0)	–	100(0)	Fe _{SPM}
		0.30(0)	0.72(0)	–	100(0)	Fe _{SPM}
M250Z	2.7	0.30(0)	0.72(0)	–	100(0)	Fe _{SPM}
M500Z	2.5	0.34(0)	0.77(0)	–	100(0)	Fe _{SPM}
M1000Z	2.7	0.35(0)	0.84(0)	–	100(0)	Fe _{SPM}

^aFe_{SPM}: superparamagnetic Fe atoms of Fe₃O₄; Fe_{Surf}: disordered surface components of Fe₃O₄; Fe_{Oct}: Fe atoms at octahedral sites of Fe₃O₄; Fe_{Tet}: tetrahedral Fe atoms; IS: isomer shift; QS: quadrupole splitting; B_{hf}: magnetic hyperfine field.

M180Z and M200Z show presence of 75 and 100% superparamagnetic nanoparticles, respectively (Table 1). Above this concentration of zeolite, the particle size was found to remain almost constant (2-3 nm) indicating saturation of adsorption.

Mössbauer parameters as a function of zeolite concentration (Figure 4) have been categorized into four different zones. In zone I the superparamagnetic component increases from 7 to 20% with introduction of 25 mg of zeolite and the net area of the tetrahedral and octahedral sites reduces to 8% along with decrease in the mean hyperfine field. These variations of the above parameters clearly indicate decrease in the net particle size distribution. In zone II the mean hyperfine field (<B_{hf}>) remains constant. Reported neutron scattering³¹ on the different particle size of Fe₃O₄ shows that the magnetic moment at the octahedral site is more affected than tetrahedral site with decrease in particle size. Therefore, the octahedral components of M25Z, M50Z and M75Z samples were possibly overlapped with surface component spectrum due to smaller particle sizes as the octahedral component may not have been resolved. The hyperfine field of octahedral sites is assumed to be 0 T for illustration. However, the area of superparamagnetic doublet increases with zeolite

concentration. Zone III shows slight increase in the area of superparamagnetic, tetrahedral and octahedral components. The <B_{hf}> values remain constant but the internal magnetic field of tetrahedral and octahedral sites increases monotonically. In Zone IV the superparamagnetic component increases with compensation to the decrease

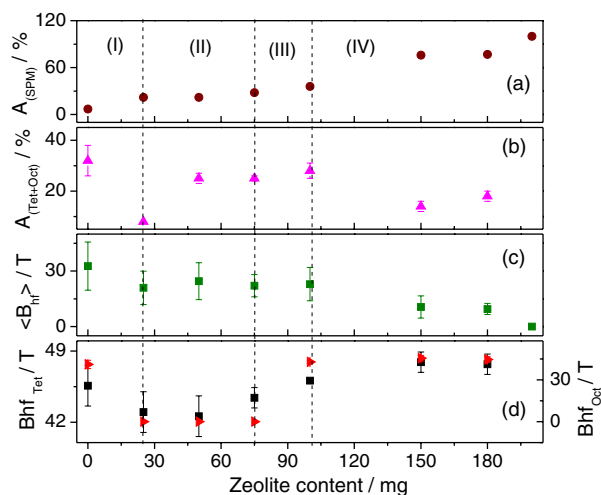


Figure 4. (a) Superparamagnetic area (A_{SPM}); (b) area of tetrahedral and octahedral ($A_{\text{Tet+Oct}}$); (c) mean hyperfine field (<B_{hf}>) and (d) hyperfine field of tetrahedral (Bhf_{Tet}) and octahedral (Bhf_{Oct}) of the sample mentioned in the x axis.

in octahedral and tetrahedral component. But the internal magnetic field experienced by the tetrahedral and octahedral Fe is larger than the pure Fe_3O_4 . Variation of all these parameters can be attributed to the decrease of area which corresponds to the larger Fe_3O_4 particles.

It is reported that zeolite attracts the ions and charged surface nanoparticles with electrostatic field from specific cationic sites.³² In low concentration, i.e., 25-75 mg of zeolite, only limited Fe ions are exchanged and small amounts of Fe_3O_4 are nucleated at the pores. In the precipitated solution nearly 20% of the precipitated iron oxide nucleated in the pores of zeolite. This reduces the concentration of free Fe_3O_4 and hence the probability of interaction reduces, which results in lower particle size from 9 to ca. 6 nm. The electrostatic field intensity emitted from the zeolite is lower than the threshold force to attract the free iron oxide nucleated in the solution.

This prevents the agglomeration of the free Fe_3O_4 on the surface of the zeolite. At higher concentrations of zeolite (100-180 mg), intense electrostatic field attracts the free Fe_3O_4 nanoparticles dispersed in the solution. These nanoparticles are attached to the surface of zeolite then agglomerates giving larger particles (low fraction) experiencing higher internal magnetic field than the pure Fe_3O_4 nanoparticles. On further increase in the zeolite concentration to 200 mg, the amount of zeolite becomes sufficient to nucleate all the Fe ions to Fe_3O_4 at the pores of the zeolite leading to sharp decrease in particle size to 3 nm. The Mössbauer spectra (Figure 5A) of the sample I0.5 consist of 51% superparamagnetic doublet contributed by the nanoparticles of less than 10 nm nucleated from the pores. The remaining 49% exhibiting sextets are from the particles of larger size nucleated on the surface.

Samples I60 and I240 show presence of 76 and 100% superparamagnetic particles, respectively. It is reported that Na^+ ions present in zeolite can be exchanged by transition metallic cations such as Mn^{2+} , Cu^{2+} , Co^{2+} and Zn^{2+} , etc.³³ As the incubation time increases, the Fe^{2+} or Fe^{3+} ions get exchanged with Na^+ ions residing even at the hexagonal connection of the tetrahedral units of AlO_4 and SiO_4 . Zeolite 13x has uniform pores with maximum diameter of 1.3 nm, with a window size of 0.8 nm (α -cage).¹⁰ Other pores that exist of 0.6 nm diameters with window size 0.22 nm are known as sodalite unit (β -cage). β -Cage is more difficult to access than α -cage and hence increasing the incubation time will provide sufficient time for the Fe^{2+} or Fe^{3+} ions to diffuse through the β -cage. Therefore, in the low incubation time (30 s), the $\text{Fe}^{2+}/\text{Fe}^{3+}$ ions exchange with the Na^+ ions is significantly less resulting in the formation of nearly 51% of the superparamagnetic Fe_3O_4 . The remaining Fe_3O_4 nanoparticles agglomerate and are adsorbed on the

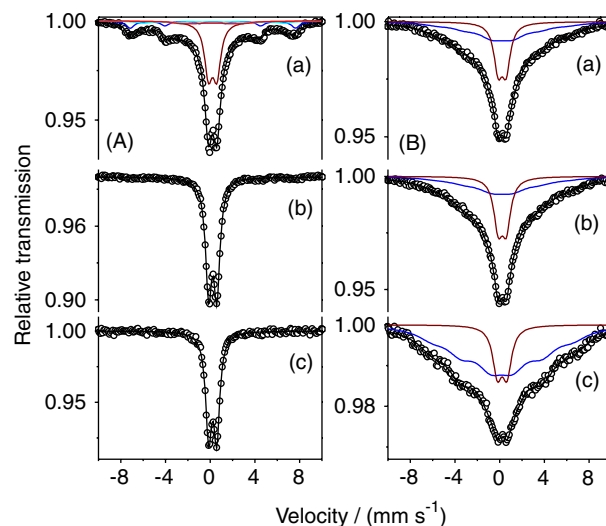


Figure 5. (A) Mössbauer spectra of Fe_3O_4 synthesized in the presence of 150 mg of zeolite with incubation time of zeolite with Fe solution of (a) 0.5 min (I0.5); (b) 60 min (I60) and (c) 240 min (I240). (B) Mössbauer spectra of the Fe_3O_4 nucleated at surface of zeolite after (a) 0.5 min (D0.5); (b) 5 min (D5) and (c) 30 min (D30) of reaction digestion time.

surface of zeolite. Increase in incubations leads to diffusion of more $\text{Fe}^{2+}/\text{Fe}^{3+}$ ions into the pores of the zeolite leading to higher fraction of superparamagnetic nanoparticles.

Figure 5B shows Mössbauer spectra of the samples D0.5, D5 and D30. The fitted parameters are shown in Table 2. The particles nucleated and attached to zeolite after 0.5 (D0.5) and 5 (D5) min of digestion time show similar superparamagnetic area (19%) and hyperfine field (ca. 36 T). This indicates that the nanoparticles are of similar particle size. Further increasing the digestion time to 30 min (D30) leads to higher mean field (38.9 T) and lower superparamagnetic doublet (7%) which is coarsening of the nanoparticles.

Zeolite 13x has maximum pore size of 1.3 nm and the iron oxide nanoparticles precipitated before addition of zeolite are larger in size than the pores of the zeolite (α -cage). Therefore, the precipitated Fe_3O_4 nanoparticles get adsorbed on the surface of the zeolite. Addition of zeolite slows down the Oswald ripening process leading to formation of larger particles. Saturation of Oswald ripening beyond the digestion time of 5 min in the chemical synthesis of Fe_3O_4 is reported.³¹ In the present case, it has been observed that beyond 5 min Oswald ripening continues leading to further agglomeration.

High temperature thermal stability of Fe_3O_4 nanoparticles has been studied in recent years for high temperature applications. Pure Fe_3O_4 nanoparticles undergo phase transition from $\gamma\text{-Fe}_2\text{O}_3$ to $\alpha\text{-Fe}_2\text{O}_3$ around 773 K.²⁶ To observe the effect of incubation time and digestion time, I0.5 and D0.5 samples were annealed at various temperatures for 2 h and the corresponding Mössbauer

Table 2. Least square fitting Mössbauer parameters of the spectra of the samples I0.5, I60, I240, D0.5, D5 and D30

Sample	IS (Fe) / (mm s ⁻¹)	QS / (mm s ⁻¹)	B _{hf} / T	Area / %	Phase ^a
I0.5	0.33(0)	0.75(0)	–	51(0)	Fe _{SPM}
	0.28(3)	0.02(5)	17.0(3)	20(0)	Fe _{Surf}
	0.41(3)	0.08(5)	39.1(4)	14(2)	Fe _{Surf}
	0.34(1)	0.01(2)	45.8(1)	15(1)	Fe _{Tet}
I60	0.34(0)	0.73(0)	–	100(0)	Fe _{SPM}
I240	0.33(0)	0.72(0)	–	100(0)	Fe _{SPM}
D0.5	0.33(1)	0.70(1)	–	19(1)	Fe _{SPM}
	0.36(1)	0.07(6)	36.2(6)	81(0)	Fe _{Surf}
D5	0.34(0)	0.72(0)	–	19(1)	Fe _{SPM}
	0.36(4)	0.08(8)	36.6(6)	81(0)	Fe _{Surf}
D30	0.33(2)	0.86(4)	–	07(1)	Fe _{SPM}
	0.32(2)	0.08(0)	38.9(4)	93(1)	Fe _{Surf}

^aFe_{SPM}: superparamagnetic Fe atoms of Fe₃O₄; Fe_{Surf}: disordered surface components of Fe₃O₄; Fe_{Tet}: tetrahedral Fe atoms; Fe_{Oct}: Fe atoms at octahedral sites of Fe₃O₄; IS: isomer shift; QS: quadrupole splitting; B_{hf}: magnetic hyperfine field.

spectra are shown in Figure 6. Mössbauer spectra of sample D0.5 annealed at 823 K shows conversion of only 3% of Fe₃O₄ to α -Fe₂O₃. This annealing temperature is 50 K higher than the complete structural transformation temperature in pure Fe₃O₄ (773 K).²⁶

In the sample I0.5, no structural transformation was observed till 873 K. The onset of structural phase transformation was observed at 923 K, where only 8% of α -Fe₂O₃ was found (Table 3). The difference in the Fe₃O₄ to α -Fe₂O₃ transition temperature and the agglomeration of the nanoparticles reflected in the area of the sextet signifies that the binding energy of Fe₃O₄ nanoparticles on the surface is comparatively low.

Figure 7 shows the thermal stability of the nanoparticles. Evolution of two endothermic peaks at 343 and 418 K is due to two different binding energies of H₂O as shown in Figure 7A. As the Fe₃O₄ nanoparticles were loaded, the second peak contributed from the dehydration of the pores disappeared as the Fe₃O₄ occupies the pores of zeolite. Figure 7B shows the increase in the endothermic peaks as zeolite concentration increases. The endothermic peak or binding energy of surface H₂O in pure Fe₃O₄ is less than the zeolite. Therefore the increase in the peak position is due to the partial binding of H₂O with zeolite surface which was masked by Fe₃O₄ that segregates outside the pore. The free surface of zeolite increases as the concentration of zeolite

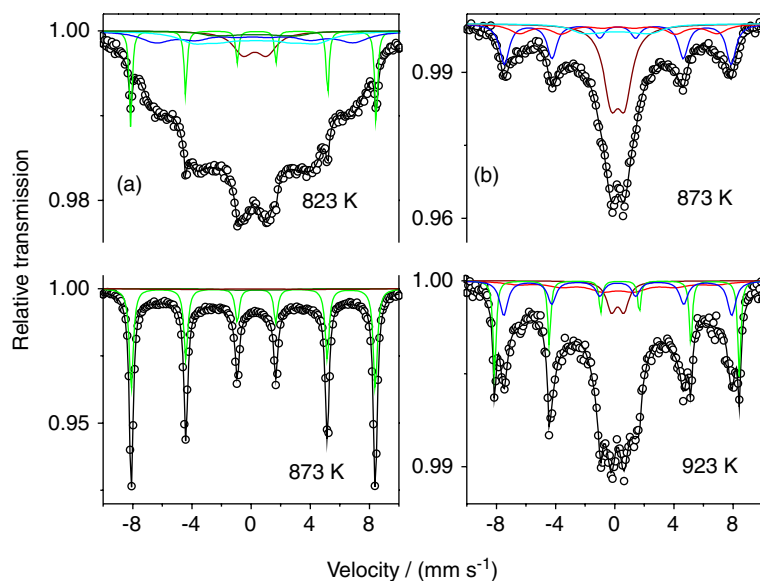
**Figure 6.** Mössbauer spectra of (a) D0.5 thermally treated at 823 and 873 K and (b) I0.5 thermally treated at 873 and 923 K.

Table 3. Least square fitting Mossbauer parameters of annealed samples D0.5 and I.05

Sample	Annealing temperature / K	IS (Fe) / (mm s ⁻¹)	QS / (mm s ⁻¹)	B _{hf} / T	Area / %	Phase ^a
D0.5	823	0.37(1)	1.60(3)	–	16(1)	Fe _{SPM}
		0.38(0)	0.02(0)	18.1(9)	07(0)	Fe _{Surf}
		0.41(0)	0.02(3)	25.7(5)	37(0)	Fe _{Surf}
		0.30(2)	0.12(4)	41.4(2)	37(0)	Fe _{Tet}
		0.38(0)	0.21(1)	51.4(0)	03(0)	Fe _{α-Fe2O3}
	873	0.40(3)	0.11(5)	10.1(2)	14(0)	Fe _{Surf}
I0.5	873	0.34(3)	0.05(4)	42.8(4)	30(1)	Fe _{Tet}
		0.37(0)	0.21(0)	51.2(0)	56(0)	Fe _{α-Fe2O3}
		0.32(0)	0.92(1)	–	38(0)	Fe _{SPM}
	923	0.30(5)	0.00(8)	15.0(0)	18(0)	Fe _{Surf}
		0.34(3)	0.01(4)	41.4(5)	22(2)	Fe _{Oct}
		0.31(1)	0.01(2)	47.5(1)	21(2)	Fe _{Tet}
		0.31(1)	0.84(3)	–	12(1)	Fe _{SPM}
		0.31(0)	0.08(6)	41.4(9)	59(1)	Fe _{Oct}
		0.32(1)	0.00(1)	47.8(0)	21(0)	Fe _{Tet}
		0.37(0)	0.21(1)	51.3(0)	08(0)	Fe _{α-Fe2O3}

^aFe_{SPM}: superparamagnetic Fe atoms of Fe₃O₄; Fe_{Surf}: disordered surface components of Fe₃O₄; Fe_{Tet}: tetrahedral Fe atoms; Fe_{α-Fe2O3}: Fe atoms associated with α-Fe₂O₃; Fe_{Oct}: Fe atoms at octahedral sites of Fe₃O₄; IS: isomer shift; QS: quadrupole splitting; B_{hf}: magnetic hyperfine field.

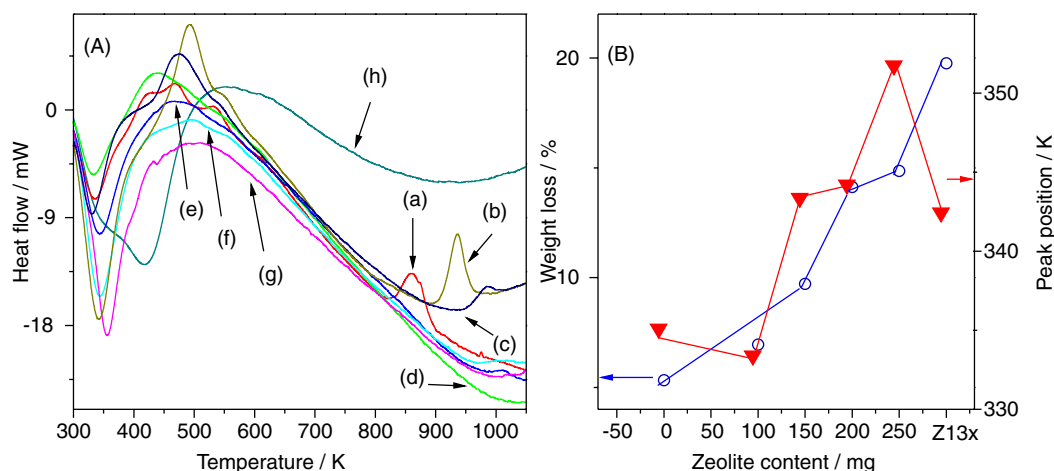


Figure 7. (A) DTA curve of the (a) M0Z, (b) M25Z, (c) M75Z, (d) M100Z, (e) M150Z, (f) M200Z, (g) M250Z and (h) zeolite 13x; (B) endothermic peaks for desorption of H₂O and total weight loss with zeolite content.

increases leading to the increase in the binding energy. In sample M250Z, the endothermic peak is greater than the first endothermic peak of pure zeolite. This is attributed to the hybridization of the binding center providing a higher energy but the pore is partially occupied providing space for the H₂O to reside in the entrance of the pore.

Figure 8a shows the variation of weight due to dehydration as a function of digestion time. The weight of the sample decreases as the digestion time increases along with shift of the endothermic peak position of dehydration towards lower temperature. This can be attributed to the

fact that nanoparticles obtained with lower digestion time enter the pores leaving the free zeolite surface for H₂O to get attached with a stronger bond. Figure 8b shows the change in weight and endothermic peak position as a function of incubation time. The mechanism of nanoparticles growth on the surface and pores of zeolite is in agreement with the TG-DTA results of the samples synthesized with different incubation times. At low incubation time (0.5 min), nearly 49% of the nanoparticles remain on the surface of the zeolite (as observed in Mössbauer spectra) which reduces the number of H₂O molecules attached to the composites

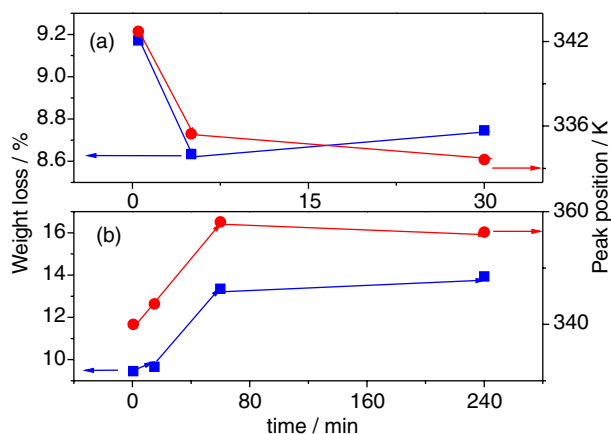


Figure 8. (a) Weight loss (■) and dehydration exothermic peak position (●) of the composites synthesized with 0.5 (D0.5), 5 (D5) and 30 (D30) min of digestion time; (b) weight loss (■) and dehydration exothermic peak position (●) of the composites synthesized with 0.5 (I0.5), 15 (I15), 60 (I60) and 240 (I240) min of incubation time.

which is reflected in the weight loss and endothermic peak. As the incubation time increases the quantity of nanoparticles in the pores increases providing more zeolite surface to interact with H_2O . Therefore, the weight loss increases from 9.4 to 13.9% as the incubation time increases from 0.5 to 240 min.

Figure 9 shows the schematic of the nucleation, growth and adsorption of Fe_3O_4 nanoparticles in the presence and absence of zeolite. Figure 9a shows that pure Fe_3O_4 seeds undergo agglomeration during synthesis resulting in larger particle size and high dipole-dipole interaction. The nanoparticle further nucleates along with structural phase transformation showing bulk hyperfine properties for annealing above 773 K. Figure 9b shows the surface-surface interaction of Fe_3O_4 nanoparticles with zeolite and depicts the nanoparticles attached to zeolite as a result of coulombic interaction in the digestion method.^{34,35} Figures 9c and 9d

show the effect of incubation time on thermal stability of Fe_3O_4 nanoparticles. Incubation provides larger fraction of nanoparticles in the pores of zeolite resulting in the increase in the superparamagnetic doublet.

The electrostatic interaction between the nanoparticles and the zeolite enhances the thermal stability. In this case a higher thermal treatment (above 773 K) is required to agglomerate the nanoparticles as well as higher activation energy for structural transformation to $\alpha\text{-Fe}_2\text{O}_3$. Surface adsorbed nanoparticles improve the thermal and chemical stability. Stability can further be increased by increasing the binding energy of the nanoparticles with zeolite. This higher binding energy can be achieved only by growing the nanoparticles from the pores of the zeolite. When the zeolite is dispersed in the precursor solution, the Na^+ ions get replaced by $\text{Fe}^{2+}/\text{Fe}^{3+}$ ions and nucleation takes place from the pores. In this case the nanoparticles get untrapped from the pores and grow only above 923 K.

Conclusions

Superparamagnetic iron oxides were synthesized in the presence of zeolite. Particle size decreases at lower concentration (25–75 mg) of zeolite. With further increase in the zeolite concentration (100–180 mg), Fe_3O_4 particles on the surface exhibits bulk magnetic properties. Nucleation on the surface in the case of digestion method maintains the crystal structure of the zeolite 13x. Amorphization takes place as the magnetite nanoparticles nucleates from the pores (incubation method). A binding energy exists within the nanoparticles and zeolite which stabilized the nanoparticles as compared to pure magnetite nanoparticles. Nanoparticles grown on the surface have lower thermal stability compared to the nanoparticles grown in the pores.

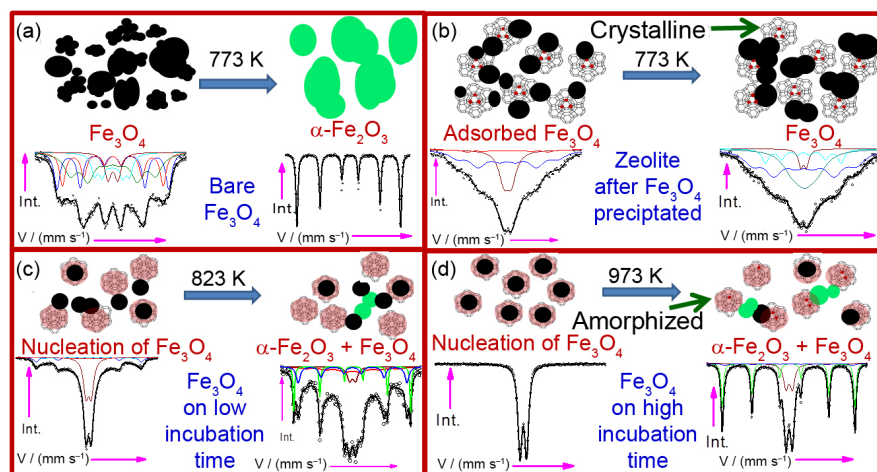


Figure 9. Schematics of the growth of nanoparticles and the thermal treatment effect on (a) pure Fe_3O_4 ; (b) Fe_3O_4 adsorbed on zeolite 13x after precipitation; (c) nucleation of Fe_3O_4 on zeolite for low incubation time and (d) nucleation of Fe_3O_4 on higher incubation of zeolite. Heat treatments were carried out in air for 2 h.

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